



HIGH SCHOOL CHEMISTRY YEAR AT A GLANCE MRS. ROSAS

TOPICS COVERED	DATES	TEKS
Scientific & Engineering Practices, Lab Safety, Measurement, Matter • Lab safety and equipment • Scientific investigations • SI units & dimensional analysis • Significant figures • Scientific notation • Density • Physical vs. chemical properties • Physical vs. chemical changes • Elements, compounds, mixtures	AUGUST 12 - SEPTEMBER 24	1A–1F (Scientific & Engineering Practices) 2A–2D (Matter & Measurements)
Atomic Structure & Periodic Table • History of atomic theory • Atomic structure • Isotopes & ions • Average atomic mass • Electron configuration • Valence electrons • Periodic trends	SEPTEMBER 28 - NOVEMBER 6	3A–3F
Chemical Bonding & Chemical Reactions • Ionic bonding • Covalent bonding • Metallic bonding • Lewis structures • Molecular geometry (VSEPR) • Naming compounds • Writing chemical formulas • Balancing equations • Types of reactions	NOVEMBER 9 - DECEMBER 18	4A–4F 5A–5D
The Mole & Stoichiometry	JANUARY 5 - FEBRUARY 12	6A–6E

• Mole concept • Molar mass • Percent composition • Empirical & molecular formulas • Mole conversions • Stoichiometry • Limiting reactants • Percent yield		
Gases & Thermochemistry • Kinetic Molecular Theory • Boyle's Law • Charles's Law • Combined Gas Law • Ideal Gas Law • Heat vs. temperature • Specific heat • Calorimetry • Endothermic & exothermic reactions	FEBRUARY 15 - APRIL 2	7A–7D 9A–9C
Solutions, Acids & Bases, Nuclear Chemistry • Solubility • Molarity • Dilutions • Acids and bases • pH scale • Neutralization • Radioactivity • Nuclear decay • Fission & fusion • Review & final exam	APRIL 5 - MAY 19	8A–8E 10A–10D