

## Topic 3: Nomenclature

1. Name or write the formula for the following ionic compounds:

a. LiCl		g. tin(II) bromide	
b. Mg(OH) <sub>2</sub>		h. potassium phosphate	
c. K <sub>3</sub> P		i. nickel(II) perchlorate	
d. Fe <sub>2</sub> O <sub>3</sub>		j. sodium hydroxide	
e. FeO		k. zinc phosphate	
f. ZnCl <sub>2</sub>		l. ammonium sulfate	

2. Name or write the formula for the following covalent compounds:

a. CO		e. nitrogen tribromide	
b. CBr <sub>4</sub>		f. tetraphosphorus decaoxide	
c. SO <sub>2</sub>		g. xenon hexafluoride	
d. N <sub>2</sub> O <sub>4</sub>		h. dicarbon tetrafluoride	

3. Name or write the formula for the following acids:

a. HCl		e. hydrobromic acid	
b. HNO <sub>3</sub>		f. hydronitric acid	
c. HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>		g. phosphoric acid	
d. H <sub>2</sub> SO <sub>4</sub>		h. hydrosulfuric acid	