

AP Chemistry Summer Assignment Directions

AP Chemistry is designed to be a second-year chemistry course. It is assumed that prior to taking AP Chemistry that you have successfully completed honors chemistry with a high A average. College Board also recommends that students have successfully completed an Algebra II course. Due to the intense nature of the course curriculum, the limited amount of time prior to the AP exam for covering the material, it is expected that you have a strong grasp of the Honors Chemistry topics, hence, the importance of this summer assignment.

If you find any of the material challenging, refer back to your notes packets from HChem, or try YouTube videos (Old School Chemistry or Khan Academy). I will **NOT** be doing a full review of the topics in the assessment packet, so you need to be solid on them. **There will be a quiz based upon the summer assignment on the SECOND day of class.** On day 3, we hit the ground running.

AP chemistry requires significant preparation and study on the part of the student. Make sure that your work skills are honed for the coming year. Expect to spend 30-45 minutes on the class every day.

If you would like to reach out to me over the summer, feel free to contact me by e-mail at lwoelfel@parrottacademy.org or via Teams msg.

I look forward to having you in AP Chemistry next year.

Have a great summer.

Mrs. Woelfel

Name: _____ Date: _____

PREP FOR AP CHEMISTRY

AP Chemistry Summer Assignment

		Due Date
Topic 1	Significant Figures	
Topic 2	Metric and Temperature Conversions	
Topic 3	Nomenclature	
Topic 4	Atomic Structure	
Topic 5	Writing and Balancing Chemical Equations	
Topic 6	Moles and Stoichiometry	
Topic 7	Graphing and Data Analysis	
Topic 8	Particulate Drawings	
Topic 9	Elements and Ions to Memorize (Prepare Flash Cards)	

Topic 1: Significant Figures

1. Determine the number of significant figures in each of the following:

a. 0.7540 _____

b. 12500 _____

c. 1000.01 _____

d. 1200 _____

e. 1.04×10^3 _____

f. 0.0080050 _____

2. Perform the following calculations and round to the appropriate number of significant figures:

a. $34.66 + 333.0$ _____

b. $1.23 + 9.66$ _____

c. $455 - 1.22$ _____

d. 18.2×1.998 _____

e. $10.2 \div 1.34$ _____

f.
$$\frac{100.23 + 59.4}{5.22}$$

3. Round each of the following numbers to three significant figures:

a. 167.789 _____

b. 0.00000445345 _____

c. 25.0545 _____

d. 3.1415926536 _____

e. 8504.0435 _____

f. 14.4355 _____

Topic 2: Metric and Temperature Conversions

1. Use dimensional analysis (factor-label method) to make the following metric conversions:

a. 3.40 m to cm _____ $\times$ _____ = _____

b. 289 cm to nm _____ $\times$ _____ $\times$ _____ = _____

c. 125145 J to kJ _____ $\times$ _____ = _____

d. 164 mg to g _____ $\times$ _____ = _____

e. 46.5 mL to L _____ $\times$ _____ = _____

2. Make the following temperature conversions:

a. 162°F to °C _____

b. 0.0 °F to K _____

c. -18 °C to K _____

d. 212 K to °C _____

e. 98.6 °F to K _____

Topic 3: Nomenclature

1. Name or write the formula for the following ionic compounds:

a. LiCl		g. tin(II) bromide	
b. Mg(OH) ₂		h. potassium phosphate	
c. K ₃ P		i. nickel(II) perchlorate	
d. Fe ₂ O ₃		j. sodium hydroxide	
e. FeO		k. zinc phosphate	
f. ZnCl ₂		l. ammonium sulfate	

2. Name or write the formula for the following covalent compounds:

a. CO		e. nitrogen tribromide	
b. CBr ₄		f. tetraphosphorus decaoxide	
c. SO ₂		g. xenon hexafluoride	
d. N ₂ O ₄		h. dicarbon tetrafluoride	

3. Name or write the formula for the following acids:

a. HCl		e. hydrobromic acid	
b. HNO ₃		f. hydronitric acid	
c. HC ₂ H ₃ O ₂		g. phosphoric acid	
d. H ₂ SO ₄		h. hydrosulfuric acid	

Topic 4: Atomic Structure

1. Determine the number of protons, neutrons and electrons in each of the following:

a. ${}^{39}_{19}\text{K}$ protons: _____ neutrons: _____ electrons: _____

b. ${}^{23}_{11}\text{Na}^{1+}$ protons: _____ neutrons: _____ electrons: _____

c. ${}^{208}_{82}\text{Pb}$ protons: _____ neutrons: _____ electrons: _____

d. ${}^{33}_{15}\text{P}^{3-}$ protons: _____ neutrons: _____ electrons: _____

2. Write the symbol for the atom that contains

a. 24 protons, 21 electrons and 24 neutrons

b. 34 protons, 45 neutrons, 34 electrons

c. 8 protons, 10 neutrons, 10 electrons

3. What experimental evidence supports these statements?

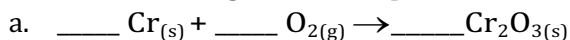
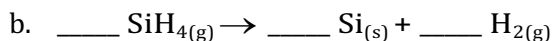
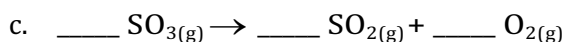
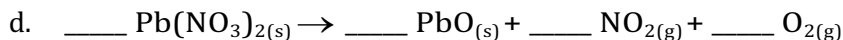
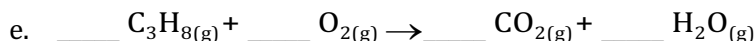
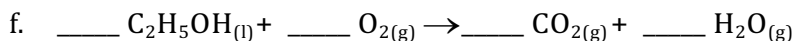
a. The nucleus of an atom is small.

b. The atom consists of both positive and negative charges.

c. The nucleus of the atom is positive.

Topic 5: Writing and Balancing Chemical Equations

1. Balance the following chemical equations:



2. Write a balanced chemical equation for each of the following reaction descriptions:

a. When solid calcium carbonate is heated, solid calcium oxide and gaseous carbon dioxide are formed.

b. Aluminum metal reacts with oxygen to form solid aluminum oxide.

c. When solid mercury(II) sulfide is heated with oxygen, liquid mercury metal and gaseous sulfur dioxide are produced.

d. When aqueous solutions of aluminum sulfate and barium chloride are mixed, solid barium sulfate and aqueous aluminum chloride are formed.

e. Solid sodium bicarbonate reacts with hydrochloric acid to produce sodium chloride, water, and carbon dioxide gas.

f. Gaseous ammonia and oxygen react to produce nitrogen monoxide gas and water

Topic 6: Moles and Stoichiometry

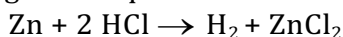
1. Vinegar is a dilute solution of acetic acid, CH_3COOH .

a. Calculate the molar mass of acetic acid.

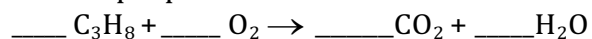
b. How many molecules of CH_3COOH are contained within 43.4 g of acetic acid?

c. How much would 0.450 moles of acetic acid weigh?

2. How many moles of hydrogen gas can be produced if 1.35 g of solid zinc reacts with excess hydrochloric acid according to the equation



3. The reaction for the combustion of propane is



a. If 20.0 g of C_3H_8 and 20.0 g of O_2 are reacted, how many moles of CO_2 can be produced?

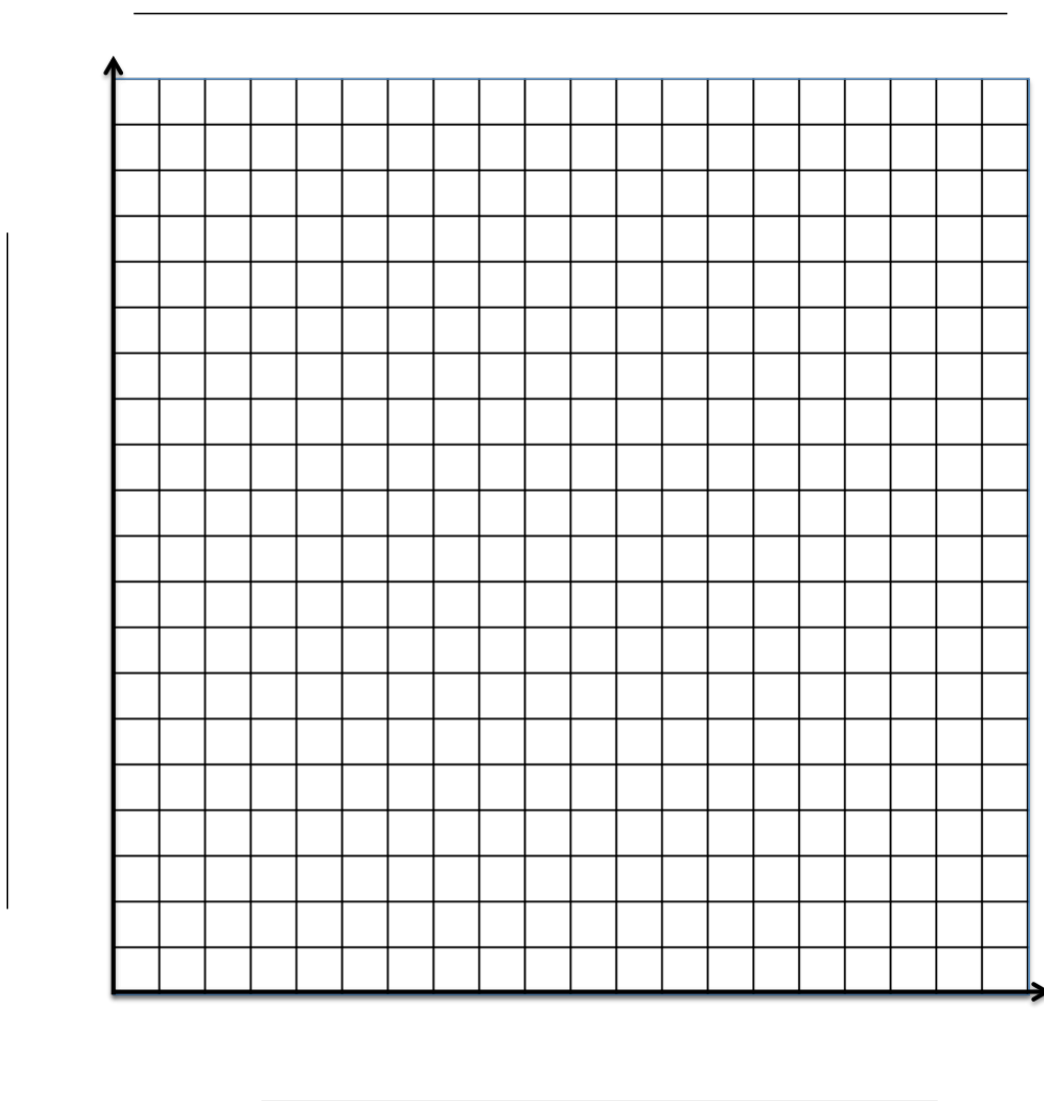
4. If 20.0 g of C_3H_8 and 80.0 g of O_2 are reacted, how many grams of CO_2 can be produced

Topic 7: Graphing and Data Analysis

1. When anhydrous calcium chloride is dissolved in water, the temperature of the system changes. A student obtains the following data when dissolving increasing amounts of CaCl_2 into 100 mL of water:

Mass of CaCl_2 dissolved, g	0.91	2.94	5.92	8.81	10.89
ΔT , $^{\circ}\text{C}$	1.8	6.6	12.8	18.9	23.2

Plot the data on the graph below. Choose an appropriate scale, and label the axes appropriately.



Refer to the graph to answer the following questions.

Independent Variable: _____

Dependent Variable: _____

Provide a descriptive title for the graph: _____

2. Describe the relationship between grams of calcium chloride salt and change in temperature in a sentence.

3. Draw a line of best fit. Determine its slope, including units.

4. Predict the change in temperature when
 - a. 4.33 g of CaCl_2 are dissolved

 - b. 9.56 g of CaCl_2 are dissolved

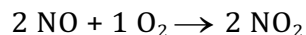
 - c. 15.4 g of CaCl_2 are dissolved

5. Predict what mass of CaCl_2 will result in
 - d. a 12.4°C change in temperature

 - e. a 44.9°C change in temperature

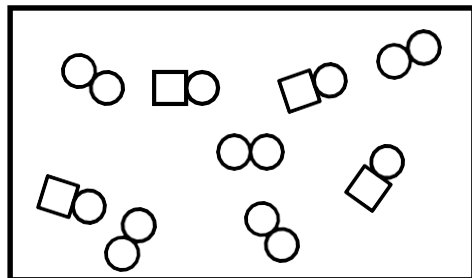
Topic 8: Particulate Diagrams

1. Consider the synthesis of nitrogen dioxide

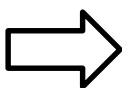


- a. In the diagram below, nitrogen atoms are represented with squares and oxygen atoms are represented with circles. Using the conservation of matter, draw what you would expect to find in the reaction vessel once the reaction is complete.

Before Reaction:



After Reaction



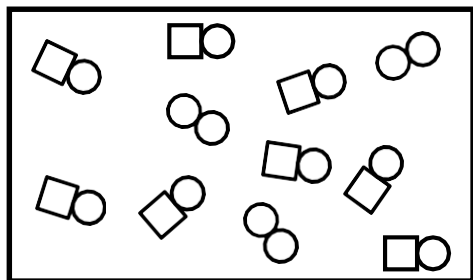
Limiting Reactant:

Excess Reactant:

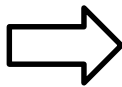
Explanation

- b. Consider the same reaction, with different starting quantities. Draw the contents of the reaction vessel after the reaction is complete.

Before Reaction:



After Reaction



Limiting Reactant:

Excess Reactant:

Explanation

Topic 9: Elements and Ions to Memorize

1. Prepare flash cards for the following elements/ions and their symbols

H	hydrogen
He	helium
Li	lithium
Be	beryllium
B	boron
C	carbon
N	nitrogen
O	oxygen
F	fluorine
Ne	neon
Na	sodium
Mg	magnesium
Al	aluminum
Si	silicon
P	phosphorus
S	sulfur
Cl	chlorine
Ar	argon
K	potassium
Ca	calcium
Cr	chromium
Mn	manganese
Fe	iron
Cu	Copper
Zn	zinc
Ag	silver
Pb	lead

$\text{C}_2\text{H}_3\text{O}_2^-$	acetate
ClO_3^-	chlorate
ClO_2^-	chlorite
CN^-	cyanide
HCO_3^-	bicarbonate
OH^-	Hydroxide
NO_3^-	nitrate
NO_2^-	nitrite
ClO_4^-	perchlorate
MnO_4^-	permanganate
SCN^-	thiocyanate
CO_3^{2-}	carbonate
CrO_4^{2-}	chromate
SO_4^{2-}	sulfate
SO_3^{2-}	sulfite
PO_4^{3-}	phosphate
PO_3^{3-}	phosphite
NH_4^+	ammonium