

Mrs. DiCrisi

**Significant digits review:**

1. Count the number of significant figures in the following measurements:

a. 2.71 g \_\_\_\_\_

b. 0.00047 kg \_\_\_\_\_

c.  $7.0 \times 10^5$  m \_\_\_\_\_

d. 1,030 L \_\_\_\_\_

e. 150 pencils \_\_\_\_\_

f. 37500 g \_\_\_\_\_

g. 0.1010 cm \_\_\_\_\_

**Scientific notation review:**

2. Express each of the following in proper scientific notation (Pay attention to sig figs and units)

a. 0.000125 m \_\_\_\_\_

b. 123,030,000 kg \_\_\_\_\_

c. 155.0 mL \_\_\_\_\_

d.  $481.9 \times 10^{-9}$  cm \_\_\_\_\_

**Add, subtract, multiply, and divide with the correct number of significant figures review:**

3. Calculate the correct answer with proper units and sig figs for each of the following:

a.  $12 \text{ g} + 0.677 \text{ g} + 86.33 \text{ g} =$  \_\_\_\_\_

b.  $(355.78 \text{ g}) / (0.056 \text{ g}) =$  \_\_\_\_\_

c.  $97.34 \text{ mL} - 34.1 \text{ mL} =$  \_\_\_\_\_

d.  $14.68 \times 5 =$  \_\_\_\_\_

4. Calculate the correct answer with proper units, sig figs, and scientific notation for each of the following:

a.  $0.14 \times (6.02 \times 10^{23}) =$  \_\_\_\_\_

b.  $\frac{(9.875 \times 10^4) - (9.795 \times 10^4)}{(9.875 \times 10^4)} \times 100\% =$  \_\_\_\_\_ (assume 100% is exact)

c.  $\frac{(3.8 \times 10^{-12} + 4.0 \times 10^{-13})}{(4 \times 10^{12} + 6.3 \times 10^{13})} =$  \_\_\_\_\_

**Dimensional analysis review:** Solve the following problems using conversions and dimensional analysis.

5. A large railroad car is filled with 1745 gallons of milk. The car springs a leak in the bottom, and milk starts dripping out at a rate of 204.84 mL/sec. If the train is traveling at a speed of 65.4 miles per hour, calculate how many miles it will travel before all the milk has drained out of the car.

(1 gal = 3.78 L, 1 mile = 5280 ft, 1 in = 2.54 cm)



**Identifying physical/chemical characteristics of matter review:**

9. Define the following terms:

- a. Solid –
- b. Liquid –
- c. Gas –
- d. Pure substance –
- e. Homogeneous mixture –
- f. Heterogeneous mixture –
- g. Chemical change –
- h. Physical change –

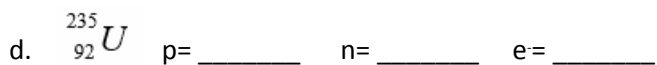
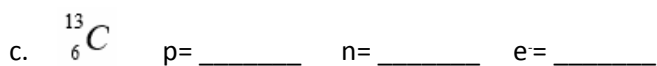
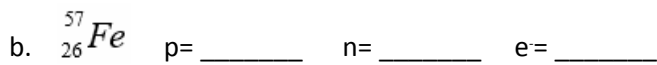
10. Identify the following as a physical property, physical change, chemical property, or chemical change:

- a. \_\_\_\_\_ Ethanol has a density of 0.697 g/mL.
- b. \_\_\_\_\_ The solution turns blue upon mixing water and food coloring.
- c. \_\_\_\_\_ Wood burns in an oven.
- d. \_\_\_\_\_ Methyl alcohol is highly flammable.
- e. \_\_\_\_\_ Ice melts in a beaker.
- f. \_\_\_\_\_ Methyl ethanoate smells like apples.
- g. \_\_\_\_\_ A car crashes into a wall.
- h. \_\_\_\_\_ Sugar dissolves in water.

**Identifying the number of protons, neutrons, and electrons in atoms and isotopes review:**

11. What number of protons and neutrons are contained in the nucleus of each of the following atoms? Assuming each atom is uncharged, what number of electrons are present?

- a.  $^{208}_{82}\text{Pb}$  p= \_\_\_\_\_ n= \_\_\_\_\_ e= \_\_\_\_\_



12. Complete the following table:

Name	Mass #	Atomic #	# of Protons	# of Neutrons	# of Electrons	Symbol
Gallium	70					
						${}_{15}^{31}\text{P}^{-3}$
Strontium-80						
						${}_{25}^{55}\text{Mn}^{+2}$

**The Law of Definite Proportions (Composition) and the Law of Multiple Proportions review:**

13. Explain:

a. Law of Definite Proportions:

b. Law of Multiple Proportions:

14. Solve the following problem:

Tin – Oxygen compound	Tin % by mass	Oxygen % by mass
Stannous oxide	88.10%	11.90%
Stannic oxide	78.70%	21.30%

Tin – Oxygen compound	Tin mass	Oxygen mass
Stannous oxide	100.0 grams	
Stannic oxide	100.0 grams	

a. Use the Law of Definite Proportions to determine the mass of oxygen needed to combine with the given masses of tin for stannous oxide and stannic oxide.

b. Does the Law of Multiple Proportions hold true in this case? Explain why or why not.

***Naming and writing formulas for ionic compounds review:***

15. Name or give the formula for the following compounds:

<u>Name</u>	<u>Formula</u>	<u>Name</u>	<u>Formula</u>
Sodium fluoride	_____	_____	K <sub>2</sub> O
Calcium phosphate	_____	_____	FeCl <sub>3</sub>
Iron (II) chloride	_____	_____	Hg <sub>2</sub> O
Sodium sulfate	_____	_____	CaCO <sub>3</sub>
Lithium phosphate	_____	_____	SO <sub>2</sub>
Calcium hydroxide	_____	_____	H <sub>2</sub> SO <sub>4</sub>
Copper (I) chloride	_____	_____	H <sub>3</sub> PO <sub>2</sub>

***Writing and balancing equations review:***

16. Write and balance the following equations:

- Iron metal reacts with oxygen gas to form solid rust, iron (III) oxide.
- Calcium metal reacts with water to produce aqueous calcium hydroxide and hydrogen gas.
- Aqueous barium hydroxide reacts with aqueous sulfuric acid to produce solid barium sulfate and water.

***Conversions associated with moles review:***

17. Solve the following problems:

- Calculate the mass of 500. atoms of iron (Fe).

- b. How many formula units are present in 87.2 grams of lead (IV) carbonate?
- c. Aspartame is an artificial sweetener that is 160 times sweeter than sucrose (table sugar) when dissolved in water. It is marketed as Nutra-Sweet. The molecular formula of aspartame is  $C_{14}H_{18}N_2O_5$ .
- Calculate the molar mass of aspartame.
  - Calculate the mass, in grams, of 1.56 mol of aspartame.
  - How many molecules are in 5.0 mg of aspartame?
  - How many atoms of nitrogen are in 1.2 g aspartame?
  - What is the mass of one molecule of aspartame?

**Percent mass calculation review:**

18. Calculate the percent by mass for each element in aspartame from the previous problem.

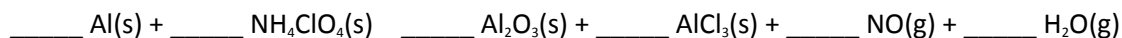
**Calculate the average atomic mass of an isotope using percent abundance review:**

19. An element consists of 1.40% of an isotope with a mass of 203.973 amu, 24.10% of an isotope with mass 205.9745 amu, 22.10% of an isotope with mass 206.9759 amu, and 52.40% of an isotope with mass 207.9766 amu. Calculate the average atomic mass and identify the element.

**Stoichiometry problems review:**

20. The reusable booster rockets of the U.S. space shuttle employ a mixture of aluminum and ammonium perchlorate for fuel. A possible reaction for this is:

- a. Balance the following reaction:



- b. If 4.0 g of aluminum reacted with 15.0 g of ammonium perchlorate, what would be the limiting reactant?

- c. How much excess of the other reactant would you have?

d. Using the above information, how much aluminum chloride would be produced in grams?

e. If you actually collected 4.18 g of aluminum chloride what would be your percent yield?

21. You add aluminum to a solution of copper (II) chloride and it reacts exothermically. Write and balance the equation below.

a. If you react 1.25 g of Al, how much copper (II) chloride do you need to add for the Al to fully react?

b. How much of each product would you collect?

22. When 125.0 g of ethylene (C<sub>2</sub>H<sub>4</sub>) burns in 60.0 grams of oxygen to give carbon dioxide and water, how many grams of CO<sub>2</sub> are formed? (Hint: balance the equation and determine limiting reactant first)

***Empirical and molecular formula calculation review:***

23. Phenol is a compound that contains 76.57% carbon, 6.43% hydrogen, and 17.0% oxygen.

a. Calculate the empirical formula.

b. If its molecular weight is 188 g/mol, what would be its molecular formula?

24. Menthol, the substance we can smell in mentholated cough drops, is composed of carbon, hydrogen, and oxygen. A 0.1005 gram sample of menthol is combusted producing 0.2829 g of CO<sub>2</sub> and 0.1159 g of H<sub>2</sub>O. What is the empirical formula for menthol? Show work.