

Single Replacement/Double Replacement Practice

Directions: Predict the products and then balance the equation.

Single Replacement

1. $\text{CaS} + \text{O}_2 \rightarrow$
2. $\text{Fe} + \text{HCl} \rightarrow$
(Hint: Fe will have a +3 charge when it bonds)
3. $\text{NaI} + \text{Br}_2 \rightarrow$
4. $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow$
5. $\text{Cl}_2 + \text{NaI} \rightarrow$
6. $\text{Fe} + \text{CuSO}_4 \rightarrow$
(Hint: Fe will have a +2 charge when it bonds)
7. $\text{F}_2 + \text{MgI}_2 \rightarrow$
8. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow$
(Hint: Zn will have a +2 charge when it bonds)
9. Sodium reacts with calcium carbonate \rightarrow
10. Fluorine reacts with sodium chloride \rightarrow

Double Replacement

1. $\text{Ca}(\text{OH})_2 + \text{HCl} \rightarrow$
2. $\text{Al}(\text{NO}_3)_3 + \text{H}_2\text{SO}_4 \rightarrow$
3. $\text{AlN} + \text{CaI}_2 \rightarrow$
4. $\text{MgBr}_2 + \text{NaCl} \rightarrow$
5. $\text{K}_2\text{S} + \text{AgI} \rightarrow$
(Hint: Ag will have a +1 charge when it bonds)
6. $\text{Al}_2(\text{SO}_4)_3 + \text{Ca}_3(\text{PO}_4)_2 \rightarrow$
7. $\text{LiCl} + \text{MgO} \rightarrow$
8. $\text{NaF} + \text{BaCl}_2 \rightarrow$
9. Sodium hydroxide reacts with sulfuric acid \rightarrow
10. ammonium chloride reacts with calcium hydroxide \rightarrow