

### Single Replacement/Double Replacement Practice

Directions: Predict the products and then balance the equation.

The ODD answers are given.

#### Single Replacement

1.  $2\text{CaS} + \text{O}_2 \rightarrow 2\text{CaO} + 2\text{S}$
2.  $\text{Fe} + \text{HCl} \rightarrow$   
(Hint: Fe will have a +3 charge when it bonds)
3.  $2\text{NaI} + \text{Br}_2 \rightarrow 2\text{NaBr} + \text{I}_2$
4.  $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow$
5.  $\text{Cl}_2 + 2\text{NaI} \rightarrow 2\text{NaCl} + \text{I}_2$
6.  $\text{Fe} + \text{CuSO}_4 \rightarrow$   
(Hint: Fe will have a +2 charge when it bonds)
7.  $\text{F}_2 + \text{MgI}_2 \rightarrow \text{MgF}_2 + \text{I}_2$
8.  $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow$   
(Hint: Zn will have a +2 charge when it bonds)
9. Sodium reacts with calcium carbonate  $\rightarrow$   
 $2\text{Na} + \text{CaCO}_3 \rightarrow \text{Na}_2\text{CO}_3 + \text{Ca}$
10. Fluorine reacts with sodium chloride  $\rightarrow$

#### Double Replacement

1.  $\text{Ca}(\text{OH})_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$
2.  $\text{Al}(\text{NO}_3)_3 + \text{H}_2\text{SO}_4 \rightarrow$
3.  $2\text{AlN} + 3\text{CaI}_2 \rightarrow 2\text{AlI}_3 + \text{Ca}_3\text{N}_2$
4.  $\text{MgBr}_2 + \text{NaCl} \rightarrow$
5.  $\text{K}_2\text{S} + 2\text{AgI} \rightarrow 2\text{KI} + \text{Ag}_2\text{S}$   
(Hint: Ag will have a +1 charge when it bonds)
6.  $\text{Al}_2(\text{SO}_4)_3 + \text{Ca}_3(\text{PO}_4)_2 \rightarrow$
7.  $2\text{LiCl} + \text{MgO} \rightarrow \text{Li}_2\text{O} + \text{MgCl}_2$
8.  $\text{NaF} + \text{BaCl}_2 \rightarrow$
9. Sodium hydroxide reacts with sulfuric acid  $\rightarrow$   
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
10. ammonium chloride reacts with calcium hydroxide  $\rightarrow$