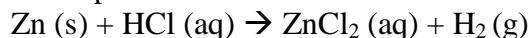


## Practice Toxins Mid-Unit Test 08-09

- \_\_\_\_\_ 1. What type of reaction is this?  $\text{Ag (s)} + \text{CuI}_2 \text{(aq)} \rightarrow \text{AgI (s)} + \text{Cu(s)}$   
(A) single displacement (B) double displacement  
(C) combination reaction (D) decomposition reaction
- \_\_\_\_\_ 2. Calcium Chloride is abbreviated  
(A) CaCl (C) Ca<sub>2</sub>Cl  
(B) CaCl<sub>2</sub> (D) Cl<sub>2</sub>Ca
- \_\_\_\_\_ 3. What is a precipitate?  
(A) a solid that forms when two solutions are mixed  
(B) rain  
(C) the moisture that forms when you are running  
(D) the formation of water in a reaction
- \_\_\_\_\_ 4. Carbon Dioxide is abbreviated  
(A) CO (B) C<sub>2</sub>O  
(C) CO<sub>2</sub> (D) O<sub>2</sub>C
- \_\_\_\_\_ 5. NaNO<sub>3</sub> is  
(A) Sodium Nitrate (B) Sodium Nitrogen  
(C) Sodium Nitrogen Oxide (D) Sodium Nitrogen Trioxide
- \_\_\_\_\_ 6. The correct chemical formula for potassium carbonate is  
(A) P<sub>2</sub>CO<sub>3</sub> (B) P(CO<sub>3</sub>)<sub>2</sub> (C) K<sub>2</sub>CO<sub>3</sub> (D) K(CO<sub>3</sub>)<sub>2</sub>
- \_\_\_\_\_ 7. NF<sub>3</sub> is  
(A) Nitrogen Fluoride (B) Nitrogen Tetrafluoride  
(C) Nitrogen Trifluoride (D) Nitrogen Fluorine
- \_\_\_\_\_ 8. Which substance is the least toxic?  
(A) chlorine (LD<sub>50</sub> = 850 mg/kg) (B) aspirin (LD<sub>50</sub> = 200 mg/kg)  
(C) cola (LD<sub>50</sub> = 140 mg/kg) (D) vitamin A (LD<sub>50</sub> = 2000 mg/kg)
- \_\_\_\_\_ 9. Br<sub>2</sub> is  
(A) Boron gas (B) Bromine gas  
(C) Bromide gas (D) DiBromine gas
- \_\_\_\_\_ 10. How many moles of potassium iodide, KI, are there in 100 grams?  
(A) 0.6 moles (B) 1660 moles (C) 3.01 x 10<sup>25</sup> moles (D) 5.0 x 10<sup>-24</sup> moles
- \_\_\_\_\_ 11. If you saw a container of NaCl (aq) in a lab, what would you see?  
(A) solid (B) liquid (C) gas
- \_\_\_\_\_ 12. In order to balance a chemical equation, you can change  
(A) the subscripts (B) the chemical formulas  
(C) the reactants and the products (D) the coefficients
- \_\_\_\_\_ 13. How many grams of sugar (LD<sub>50</sub> = 30 g/kg) would be lethal to a 175-pound man (2.2 lbs = 1kg)?  
(A) 0.13 g (B) 2.65 g (C) 11550 g (D) 2386 g
- \_\_\_\_\_ 14. Calculate the molar mass of ammonium phosphate, (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>.  
(A) 96.0 g/mole (B) 113.0 g/mole (C) 242.0 g/mole (D) 121.0 g/mole
- \_\_\_\_\_ 15. What would you expect to see if you performed the following chemical reaction?  
 $2 \text{H}_2\text{O}_2 \text{(aq)} \rightarrow \text{H}_2\text{O (l)} + \text{O}_2 \text{(g)}$   
(A) bubbles forming in a liquid (B) just a liquid  
(C) solid forming in a liquid (D) only solid forming
- \_\_\_\_\_ 16.  $\text{CO}_2 \text{(s)} \rightarrow \text{CO}_2 \text{(g)}$  is an example of  
(A) a chemical change (B) a physical change  
(C) a combination reaction (D) a decomposition reaction

\_\_\_\_\_ 17. What is the correct description of this reaction?



- (A) Hydrogen gas reacts with a solution of zinc chloride to produce solid zinc metal in a solution of hydrochloric acid.
- (B) A solution of zinc metal reacts with solid hydrochloric acid to produce a solution of zinc chloride and hydrogen gas.
- (C) Hydrogen gas reacts with solid zinc chloride to produce solid zinc metal in a solution of hydrochloric acid.
- (D) Solid zinc reacts with a solution of hydrochloric acid to produce a solution of zinc chloride and hydrogen gas.

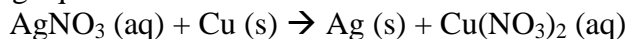
\_\_\_\_\_ 18. When the equation  $\text{Fe}_2\text{O}_3 + \text{H}_2 \rightarrow \text{Fe} + \text{H}_2\text{O}$  is balanced, Fe has a coefficient of

- (A) 6 (B) 3 (C) 2 (D) 1

\_\_\_\_\_ 19. What type of reaction is this:  $\text{AgNO}_3 \text{ (aq)} + \text{Cu (s)} \rightarrow \text{Ag (s)} + \text{Cu(NO}_3)_2 \text{ (aq)}$ ?

- (A) single displacement (B) double displacement (C) combination (D) decomposition

\_\_\_\_\_ 20. Balance the following equation. What will the coefficient be for Silver Nitrate?



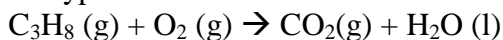
- (A) 1 (B) 2  
(C) 3 (D) 4

Short Answer

21. Consider the following reaction. Water ( $\text{H}_2\text{O}$ ) is produced from the reaction of oxygen gas ( $\text{O}_2$ ) with hydrogen gas ( $\text{H}_2$ ).

- a. Write a balanced chemical equation for the reaction.  
b. Is the reaction a physical change or a chemical change?

22. What type of reaction is shown below?



23. State the reaction type and balance the following equation:  $\text{Pb (s)} + \text{I}_2 \text{ (g)} \rightarrow \text{PbI}_2 \text{ (s)}$

24. State the names of each compound:

- a.  $\text{AlCl}_3$   
b.  $\text{Ca(OH)}_2$

25. Write the correct formulas for the following compounds:

- a. Potassium Sulfide  
b. Dinitrogen tetrafluoride

Answers:

- 1 a 8 d 15 a 25.  
2. b 9 b 16 b  
3 a 10 a 17 d  
4. c 11 b 18 c  
5. a 12 d 19 a  
6 c 13 d 20 b  
7 c 14 a 21 a  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  b Chemical Change  
22 combustion  
23 synthesis (or combination); already balanced  
24 Double replacement;  $2\text{AlCl}_3 + 3\text{Ca(OH)}_2 \rightarrow 2\text{Al(OH)}_3 + 3\text{CaCl}_2$   
25 a.  $\text{K}_2\text{S}$  b.  $\text{N}_2\text{F}_4$