

**Marietta City Schools**  
**2024–2025 District Unit Planner**

<b>IB Chemistry SL Year 1</b>		<b>Subject Group and Course</b>	Group 4 - Chemistry		
<b>Course Part and Topic</b>	<b>UNIT 1 - ATOMS AND PERIODICITY</b> Structure 1.2 - The nuclear atom Structure 1.3 - Electron configurations Structure 3.1 - The periodic table: Classification of elements	<b>SL or HL / Year 1 or 2</b>	SL / Year 1	<b>Dates</b>	08/01 - 09/11
<b>Text(s)</b>		<b>DP Assessment(s) for Unit</b>			
<ul style="list-style-type: none"> <li>Chemistry for the IB Diploma Third Edition, Hodder Education</li> </ul>		<ul style="list-style-type: none"> <li>Unit 01 Summative Assessment - <i>Paper 1 and 2 questions modeled after the real IB Exam Papers (2025 syllabus)</i></li> </ul>			

***INQUIRY: establishing the purpose of the unit***

**Transfer Goals**

List here one to three big, overarching, long-term goals for this unit. Transfer goals are the major goals that ask students to “transfer” or apply their knowledge, skills, and concepts at the end of the unit under new/different circumstances, and on their own without scaffolding from the teacher.

Phenomenon: Isotopes are used in medical imaging to diagnose and monitor a wide variety of conditions through the interactions that they have with electrons in the human body.

Statement of Inquiry: The fundamental principles that shape the behavior and properties of elements enable scientists to develop and use predictive models across scientific disciplines.

Goals:

- Students can** explain how the nuclei of atoms differ.
- Students can** model the energy states of electrons in atoms.
- Students can** use the periodic table to predict patterns and trends in the properties of elements.

***ACTION: teaching and learning through inquiry***

Content / Skills / Concepts - Essential Understandings	Learning Process
<p><b>Structure 1.2.1</b>  <b>Atoms contain a positively charged, dense nucleus composed of protons and neutrons (nucleons). Negatively charged electrons occupy the space outside the nucleus.</b>  <i>Use the nuclear symbol <math>{}^A_ZX</math> to deduce the number of protons, neutrons and electrons in atoms and ions.</i></p> <ul style="list-style-type: none"> <li>Relative masses and charges of the subatomic particles should be known; actual values are given in the data booklet. The mass of the electron can be considered negligible.</li> </ul> <p><b>Structure 1.2.2</b>  <b>Isotopes are atoms of the same element with different numbers of neutrons.</b>  <i>Perform calculations involving non-integer relative atomic masses and abundance of isotopes from given data.</i></p> <ul style="list-style-type: none"> <li>Differences in the physical properties of isotopes should be understood.</li> <li>Specific examples of isotopes need not be learned.</li> </ul> <p><b>Structure 1.3.1</b>  <b>Emission spectra are produced by atoms emitting photons when electrons in excited states return to lower energy levels.</b>  <i>Qualitatively describe the relationship between colour, wavelength, frequency and energy across the electromagnetic spectrum.</i>  <i>Distinguish between a continuous and a line spectrum.</i></p> <ul style="list-style-type: none"> <li>Details of the electromagnetic spectrum are given in the data booklet.</li> </ul>	<p><i>Check the boxes for any pedagogical approaches used during the unit. Aim for a variety of approaches to help facilitate learning.</i></p> <p>Learning experiences and strategies/planning for self-supporting learning:</p> <ul style="list-style-type: none"> <li><input checked="" type="checkbox"/> Lecture</li> <li><input type="checkbox"/> Socratic seminar</li> <li><input checked="" type="checkbox"/> Small group/pair work</li> <li><input checked="" type="checkbox"/> PowerPoint lecture/notes</li> <li><input checked="" type="checkbox"/> Individual presentations</li> <li><input checked="" type="checkbox"/> Group presentations</li> <li><input checked="" type="checkbox"/> Student lecture/leading</li> <li><input type="checkbox"/> Interdisciplinary learning</li> </ul> <p>Details:</p> <p><i>Students will learn through a combination of presentations, small group work, practice problems, and lab work.</i></p> <ul style="list-style-type: none"> <li><input checked="" type="checkbox"/> Other(s): <i>practice problems, lab work</i></li> </ul>
<p><b>Structure 1.3.2</b>  <b>The line emission spectrum of hydrogen provides evidence for the existence of electrons in discrete energy levels, which converge at higher energies.</b>  <i>Describe the emission spectrum of the hydrogen atom, including the relationships between the lines and energy transitions to the first, second and third energy levels.</i></p> <ul style="list-style-type: none"> <li>The names of the different series in the hydrogen emission spectrum will not be assessed.</li> </ul>	<p><b>Formative assessment(s):</b></p> <p><i>Short closer quizzes for each lesson</i>  <i>Practice with Tools and Inquiries</i>  <i>Daily formative checks</i></p>
<p><b>Structure 1.3.3</b>  <b>The main energy level is given an integer number, n, and can hold a maximum of <math>2n^2</math> electrons.</b>  <i>Deduce the maximum number of electrons that can occupy each energy level.</i></p>	<p><b>Summative assessments:</b></p>

<p><b>Structure 1.3.4</b> A more detailed model of the atom describes the division of the main energy level into s, p, d and f sublevels of successively higher energies. <i>Recognize the shape and orientation of an s atomic orbital and the three p atomic orbitals.</i></p>	<p>Content Exam - <i>Items to gauge content mastery</i></p> <p>DP Assessment - <i>Paper 1 and 2 questions modeled after the real IB Exam Papers (2025 syllabus)</i></p> <p>Laboratory Assignment - <i>assessing Tools and Inquiries practiced in the Unit</i></p>
<p><b>Structure 1.3.5</b> Each orbital has a defined energy state for a given electron configuration and chemical environment, and can hold two electrons of opposite spin. Sublevels contain a fixed number of orbitals, regions of space where there is a high probability of finding an electron. <i>Apply the Aufbau principle, Hund's rule and the Pauli exclusion principle to deduce electron configurations for atoms and ions up to Z = 36.</i></p> <ul style="list-style-type: none"> <li>• Full electron configurations and condensed electron configurations using the noble gas core should be covered.</li> <li>• Orbital diagrams, i.e. arrow-in-box diagrams, should be used to represent the filling and relative energy of orbitals.</li> <li>• The electron configurations of Cr and Cu as exceptions should be covered.</li> </ul> <p><b>Structure 3.1.1</b> The periodic table consists of periods, groups, and blocks. <i>Identify the positions of metals, metalloids, and non-metals in the periodic table.</i></p> <ul style="list-style-type: none"> <li>• The four blocks associated with the sublevels s, p, d, f should be recognized.</li> <li>• A copy of the periodic table is available in the data booklet.</li> </ul>	<p><b>Differentiation:</b></p> <ul style="list-style-type: none"> <li><input checked="" type="checkbox"/> Affirm identity - build self-esteem</li> <li><input checked="" type="checkbox"/> Value prior knowledge</li> <li><input checked="" type="checkbox"/> Scaffold learning</li> <li><input checked="" type="checkbox"/> Extend learning</li> </ul> <p>Details:</p> <ul style="list-style-type: none"> <li>• SWD/504 – Accommodations Provided</li> <li>• ELL – Reading &amp; Vocabulary Support</li> <li>• Intervention Support</li> <li>• Extensions – Enrichment Tasks and Project</li> </ul>
<p><b>Structure 3.1.2</b> The period number shows the outer energy level that is occupied by electrons. Elements in a group have a common number of valence electrons. <i>Deduce the electron configuration of an atom up to Z = 36 from the element's position in the periodic table and vice versa.</i></p> <ul style="list-style-type: none"> <li>• Groups are numbers from 1 to 18.</li> <li>• The classifications “alkali metals”, “halogens”, “transition elements”, and “noble gases” should be known.</li> </ul> <p><b>Structure 3.1.3</b> Periodicity refers to trends in properties of elements across a period and down a group. <i>Explain the periodicity of atomic radius, ionic radius, ionization energy, electron affinity, and electronegativity.</i></p>	<p><b>Tools and Inquiries:</b></p> <p><b>Structure 1.3.2</b></p> <ul style="list-style-type: none"> <li>• Inquiry 2—In the study of emission spectra from gaseous elements and of light, what qualitative and quantitative data can be collected from instruments such as gas discharge tubes and prisms?</li> </ul> <p><b>Structure 3.1.4</b></p> <ul style="list-style-type: none"> <li>• Inquiry 2, Tool 2 - Why are simulations often used in exploring the trends in chemical reactivity of group 1 and group 17 elements?</li> </ul>

**Structure 3.1.4**

**Trends in properties of elements down a group include the increasing metallic character of group 1 elements and decreasing non-metallic character of group 17 elements.**

*Describe and explain the reactions of group 1 metals with water, and of group 17 elements with halide ions.*

**Structure 3.1.5**

**Metallic and non-metallic properties show a continuum. This includes the trend from basic metal oxides through amphoteric to acidic non-metal oxides.**

*Deduce equations from the reactions with water of the oxides of group 1 and group 2 metals, carbon and sulfur.*

- Include acid rain caused by gaseous non-metal oxides, and ocean acidification caused by increasing CO<sub>2</sub> levels.

**Structure 3.1.6**

**The oxidation state is a number assigned to an atom to show the number of electrons transferred in forming a bond. It is the charge that atom would have if the compound were composed of ions.**

*Deduce the oxidation states of an atom in an ion or a compound.*

*Explain why the oxidation state of an element is zero.*

- Oxidation states are shown with a + or - sign followed by the Arabic symbol for the number, e.g. +2, -1. Examples should include hydrogen in metal hydrides (-1) and oxygen in peroxides (-1).
- The terms “oxidation number” and “oxidation state” are often used interchangeably, and either term is acceptable in assessment.
- Naming conventions for oxyanions use oxidation numbers shown with Roman numerals, but generic names persist and are acceptable. Examples include NO<sub>3</sub><sup>-</sup> nitrate, NO<sub>2</sub><sup>-</sup> nitrite, SO<sub>4</sub><sup>2-</sup> sulfate, SO<sub>3</sub><sup>2-</sup> sulfite.

**Approaches to Learning (ATL)**

*Check the boxes for any explicit approaches to learning connections made during the unit. For more information on ATL, please see [the guide](#).*

Thinking

Social

Communication

Self-management

Research

Details:

- *Students will be continuously challenged to develop higher-order thinking skills as they take prior knowledge, combine it with new content, and synthesize new understandings and connections.*
- *Students will build social groups through group work and intentional reflection activities.*
- *Students will communicate their findings to their peers in the form of small-group presentations.*
- *Students will continue to work on self-management and organization skills.*
- *Students will complete background research to develop and extend their learning.*

Language and Learning	TOK Connections	CAS Connections
<p><i>Check the boxes for any explicit language and learning connections made during the unit. For more information on the IB's approach to language and learning, please see <a href="#">the guide</a>.</i></p>	<p><i>Check the boxes for any explicit TOK connections made during the unit</i></p>	<p><i>Check the boxes for any explicit CAS connections. If you check any of the boxes, provide a brief note in the "details" section explaining how students engaged in CAS for this unit.</i></p>

<ul style="list-style-type: none"> <li><input checked="" type="checkbox"/> Activating background knowledge</li> <li><input checked="" type="checkbox"/> Scaffolding for new learning</li> <li><input checked="" type="checkbox"/> Acquisition of new learning through practice</li> <li><input checked="" type="checkbox"/> Demonstrating proficiency</li> </ul> <p>Details:</p> <ul style="list-style-type: none"> <li>• Content and vocabulary introduced in previous science courses will be used in this unit.</li> <li>• Students will use many of the concepts from this unit in future units throughout the two-year course.</li> <li>• Students will acquire new vocabulary.</li> <li>• Students will continually demonstrate proficiency with chemistry vocabulary in class discussions and group work.</li> </ul>	<ul style="list-style-type: none"> <li><input type="checkbox"/> Personal and shared knowledge</li> <li><input checked="" type="checkbox"/> Ways of knowing</li> <li><input type="checkbox"/> Areas of knowledge</li> <li><input type="checkbox"/> The knowledge framework</li> </ul> <p>Details:</p> <ul style="list-style-type: none"> <li>• TOK knowledge questions will be included as discussion options for each lesson.</li> </ul>	<ul style="list-style-type: none"> <li><input checked="" type="checkbox"/> Creativity</li> <li><input type="checkbox"/> Activity</li> <li><input type="checkbox"/> Service</li> </ul> <p>Details:</p> <ul style="list-style-type: none"> <li>• Students will be encouraged to consider the creativity involved in scientific experimentation. Students can explore alternative ways (visual, for example) to express and explain this creativity to others.</li> </ul>
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### Resources

List and attach (if applicable) any resources used in this unit

Resources for 2025 Syllabus:

- Chemistry for the IB Diploma Third Edition, Hodder Education
- [IB Chemistry Guide First Assessment 2025](#)
- InThinking IB subject site for Chemistry
- IB Chemistry Schoology Course

### **REFLECTION: considering the planning, process, and impact of the inquiry**

<b>What worked well</b> <i>List the portions of the unit (content, assessment, planning) that were successful</i>	<b>What didn't work well</b> <i>List the portions of the unit (content, assessment, planning) that were not as successful as hoped</i>	<b>Notes / Changes / Suggestions</b> <i>List any notes, suggestions, or considerations for the future teaching of this unit</i>

