

1. Endothermic reactions vs exothermic reactions. Difference? Similarity?
2. What are units for energy?
3. The specific heat of metal X is 1.5 cal/gdegC . What does this mean? Use it in calculations
4. Food calories are really _____
5. The specific heat of water compared to metals is _____ this means _____
6. What will happen to a 2 liter balloon if you increase the temperature and decrease the pressure?
7. What is the volume of 2 moles CO₂ at 200K and 3atm?
8. When temp is held constant P and T are _____ proportional. P and V are _____ V and T are _____
9. Names associated with gas laws and what they found.
10. What is normal atmospheric pressure at standard conditions? Give all
11. We made a PV graph in class, describe it and why it has the shape it does.
12. Which gases will DIFFUSE fastest ? _____ slowest? _____
13. Intermolecular forces vs intramolecular forces. Summarize similarities, differences. Molecules held together with which one? Surface tension due to which one?
14. Water has lots of surface tension and forms many H-bonds. What does this result in?
15. The process of changing from the solid to liquid phase is called what? G to L? L to G? S to G?
16. The heat of fusion of water is _____ this means it takes _____ how about Hv?
17. Weak intermolecular forces of attraction are likely to cause SOLIDS, LIQUIDS OR GASES?
18. How much heat does it take to melt 3 grams of ice?
19. how much heat does it take to boil 3 grams of water?
20. CO₂ dissolves best in _____ temp water. Sugar dissolves best in _____ temp water.
21. What is boiling point elevation? _____ Freezing point depression? _____
22. A tincture is _____
23. We usually use molarity for solution concentration, when do we use molality?
24. Saturated means _____ unsaturated? _____
25. What is the dissociation factor for sugar? NaCl? CaCl₂
26. Who is your favorite chemistry teacher.
27. What is your best guess for the freezing point of pop?
28. Bases are _____ acids are _____
29. If the $[H^+] = 10^{-3}$ then pH = _____ pOH = _____
30. A strong acid dissociates _____
31. The molarity of a solution is the measure of the number of _____ in each _____.
32. If the pH of a solution is 3 then the $[H^+] =$ _____ A) 10^{-3} B) 10^{-11} C) 3 D) 11
33. Name some acid base indicators.
34. HCl + calcium carbonate yields _____ + _____ + _____
35. The dissociation of Al(OH)₃ → _____ + _____
36. True (A) or False(B) - The label on a bottle of HCl says 12M. This is strong and dilute.
37. What is the chemical formula for phosphoric acid?
38. A pOH of 7 tells you that the solution is _____
39. What is the name of this organic compound? CH₃CH₂CH₂(OH)
40. Review organic chemistry – ALL of it.
41. Yes
42. Uh huh
43. Use the solubility chart in your book page 414
44. Vinegar has many names and different ways to draw it.
48. Atmospheric air pressure last week was 1.1 atm. What is this in torr? kPa?
49. Circle one I PREFER LABS _____ I PREFER CALCULATIONS IN CLASS _____
50. Are you taking AP chemistry next year?