This checklist includes <u>all units</u> for the Edexcel GCSE (9-1) Combined Science Course. Please ask your teachers about which units have been covered in Years 9, 10 and 11.

You can also access the individual unit checklists from www.revise4science.weebly.com

BIOLOGY

CB1 Key Concepts in Biology (Paper 1: Biology 1 and Paper 4: Biology 2)

CB1a Microscopes

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall what an electron microscope is.			
5 th	Recall what is meant by an instrument's resolution.			
5 ^{ch}	Explain why some cell structures can be seen with an electron microscope but not with a light microscope.			
7 th	Calculate total magnification using an equation.			
7 th	Calculate sizes using magnifications.			
5 th	Interpret the SI prefixes milli-, micro-, nano- and pico			

CB1b Plant and animal cells

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ^{ch}	Identify the parts of plant and animal cells.			
5 th	Recall the parts of plant and animal cells.			
5 th	Make drawings of plant and animal cells using a light microscope and identify their parts.			
6 th	Describe the functions of the sub-cellular structures commonly found in eukaryotic cells (nucleus, cell membrane, cell wall, chloroplasts, mitochondria and ribosomes).			
6 th	Estimate sizes using microscope fields of view.			
6 th	Estimate sizes using scale bars.			

CB1c Specialised cells

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe how sperm cells are adapted to their function.			
6 th	Describe how egg cells are adapted to their function.			
6 th	Describe how ciliated epithelial cells are adapted to their function.			
7 th	Draw conclusions about a cell's function from its adaptations.			

CB1d Inside bacteria

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Identify the common parts of bacteria.			
5 th	Describe the functions of common parts of bacteria.			
6 th	Describe why bacteria are classified as being prokaryotic.			
6 th	Change numbers to and from standard form.			
8 th	Compare eukaryotic and prokaryotic cells.			

CB1e Enzymes and nutrition

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	State that enzymes are proteins.			
6 th	Give examples of enzymes and where they are found in the human body and in other species.			
6 th	Recall the subunits from which carbohydrates, proteins and lipids are formed (sugars, amino acids, fatty acids and glycerol).			
6 th	Describe what enzymes do (catalyse the synthesis and breakdown of substances, such as carbohydrates, proteins and lipids, by speeding up the rate of reaction).			
6 th	Define an enzyme as a biological catalyst.			
7 th	Explain why catalysis by enzymes is important for life processes (because reactions happen much faster).			

CB1f Enzyme action

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State what enzyme specificity means.			
7 th	State that an enzyme's action is due to its active site.			
7 th	Describe the role of the active site in enzyme function (including specificity).			
9 th	Use the lock-and-key model to develop explanations for enzyme activity.			
8 th	Explain why enzymes have a particular shape, as a result of the sequence of amino acids in the chain.			
9 th	Explain how enzymes become denatured.			

CB1g Enzyme activity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Describe the effect of temperature on enzyme activity.			
8 th	Describe the effect of substrate concentration on enzyme activity.			
8 th	Describe the effect of pH on enzyme activity.			
8 th	Explain what is meant by the optimum pH/temperature of an enzyme.			
9th	Calculate the rate of enzyme activity from experimental data.			
9th	Explain why temperature, substrate concentration and pH affect enzyme activity.			

CB1h Transporting substances

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State that substances are transported by diffusion, osmosis and active transport.			
7 th	Describe how substances are transported by active transport (including the need for energy).			
6 th	Explain how substances are transported by diffusion.			
9 th	Explain how substances are transported by osmosis.			
9 th	Explain the effects of osmosis on cells and tissues.			
8 th	Investigate osmosis in potatoes.			
9 th	Calculate percentage gain and loss of mass in osmosis.			

CB2 Cells and control (Paper 1: Biology 1)

CB2a Mitosis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	List the names and order of the stages of the cell cycle, including mitosis.			
8 th	Describe what happens in each stage of the cell cycle, including mitosis.			
7 th	Describe why mitosis is important for an organism. (growth, repair, asexual reproduction)			
9th	Explain why organisms may rely on asexual reproduction.			
7 th	Describe how mitosis produces genetically identical, diploid cells.			
7 th	Describe how cancers grow.			

CB2b Growth in animals

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Define growth in animals as an increase in cell number and size.			
5 th	Give examples of specialised animal cells.			
6 th	Describe how structure of specialised animal cells is related to their function.			
7 th	Explain why cell differentiation is important in the development of specialised cells.			
8 th	Use percentile growth curves to interpret growth in children.			

CB2c Growth in plants

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Describe the stages of growth in plants (cell division/mitosis, elongation, differentiation).			
5 th	Give examples of specialised plant cells.			
6 th	Describe how the structures of specialised plant cells are related to their functions.			
7 th	Explain why cell differentiation is important in the development of specialised cells in plants.			

CB2d Stem cells

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe where stem cells are found.			
7 th	Describe the function of stem cells in plants and animals.			
9 th	Compare embryonic and adult stem cells in animals.			
7 th	Give examples of where stem cells may be used in medicine.			
8 th	Identify benefits and risks of using stem cells in medicine.			
10 ^{ch}	Evaluate the use of stem cells in medicine (by comparing their benefits and risks).			

CB2e The nervous system

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	List the parts of the nervous system.			
4 th	Describe how the nervous system detects stimuli.			
7 th	Describe the structure of sensory neurones.			
7 th	Describe the routes that impulses take to and from the brain.			
8 th	Explain how sensory neurones are adapted to their functions (including the myelin sheath).			

CB2f Neurotransmission speeds

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how the nervous system responds to stimuli.			
7 th	Describe the structures of motor neurones and relay neurones.			
8 th	Explain how motor neurones are adapted to their functions.			
9 th	Explain the action and function of synapses.			
9 th	Explain how the structure of the reflex arc allows a faster response.			
8 th	Describe the structure and function of the reflex arc.			

CB3 Genetics (Paper 1: Biology 1)

CB3a Meiosis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Recall that gametes are produced by meiosis.			
8 th	Describe what happens in meiosis. [without details of the stages]			
8 th	Explain why haploid gametes are needed for sexual reproduction.			
6 th	Recall what an organism's genome is.			
6 th	Describe where genes are found.			
6 th	Recall the function of genes.			

CB3b DNA

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall where DNA is found in a eukaryotic cell.			
7 th	Name the bases in DNA.			
7 th	Recall the pairing of bases in DNA.			
7 th	Describe how DNA strands are held together.			
8 th	Describe the overall structure of DNA.			
7 th	Describe how DNA can be extracted from fruit.			

CB3c Alleles

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe the difference between a gene and an allele.			
8 th	Explain the effects of alleles on inherited characteristics.			
7 th	Describe the relationship between a genotype and a phenotype.			
7 th	Identify homozygous and heterozygous genotypes.			
9 th	Use genetic diagrams to work out possible combinations of alleles in the offspring of parents.			
9th	Explain why the effects of some alleles in an organism's genotype are not seen in its phenotype.			

CB3d Inheritance

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Use Punnett squares to work out possible combinations of alleles in the offspring of parents.			
9 th	Interpret family pedigree charts to work out possible inherited genotypes and phenotypes.			
6 th	Describe how sex is determined in humans.			
9 th	Calculate ratios of phenotypes (controlled by alleles of a single gene) when organisms are crossed.			
9 th	Calculate probabilities of certain phenotypes occurring when organisms are crossed.			

CB3e Gene mutation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Give examples of characteristics controlled by multiple genes.			
6 th	Define the term mutation.			
6 th	Describe some potential applications of mapping human genomes.			
9 th	Explain how a mutation can cause variation (limited to changes in the protein formed, which can affect processes in which that protein is needed).			
7 th	Give examples of mutations in human genes that affect the phenotype, and examples of those that have little or no obvious effect.			
8 th	Explain why many mutations have no effect on the phenotype.			

CB3f Variation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Distinguish between genetic variation and environmental variation.			
5 th	Distinguish between continuous and discontinuous variation.			
6 th	Describe the causes of genetic variation (mutation and sexual reproduction).			
6 th	Describe the causes of environmental variation (differences in the environment, acquired characteristics).			
7 th	Analyse the contribution of genes and environment to the variation in a characteristic.			

CB4 Natural Selection and Genetic Modification (Paper 1: Biology 1)

CB4a Evidence for human evolution

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Define 'evolution'.			
5 ^{ch}	Recognise binomial species names.			
7 th	Explain how evidence from fossils and stone tools supports current ideas about human evolution.			
5 ^{ch}	Recall how stone tools are dated from their environment.			
6 th	Describe how stone tools created by human- like species have developed over time.			
6 th	Describe the fossil evidence for human-like species that lived 4.4, 3.2 and 1.6 million years ago.			

CB4b Darwin's theory

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall the cause of genetic variation.			
5 ch	Describe how adaptations allow organisms to survive.			
8 th	Explain how natural selection allows some members of a species to survive better than others when conditions change.			
9th	Explain how natural selection can lead to the evolution of a new species.			
10 th	Explain how the development of resistance in organisms supports Darwin's theory.			

CB4c Classification

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ^{ch}	Describe how organisms are classified into smaller and smaller groups (based on their characteristics).			
6 th	Identify genus and species from a binomial name.			
6 th	Identify an organism as a member of one of the five kingdoms.			
7 th	Describe what genetic analysis is.			
9 th	Explain why biologists often now classify organisms into three domains.			

CB4d Breeds and varieties

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe why new breeds and varieties are created.			
7 th	Describe what is meant by a 'genetically modified organism'.			
8 th	Describe how selective breeding is carried out.			
10 ^{ch}	Explain the impact of selective breeding on domesticated plants and animals.			

CB4e Genes in agriculture and medicine

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9th	Describe the main stages of genetic engineering.			
7 th	Recall some uses of selectively bred organisms (in agriculture).			
8 th	Recall some uses of genetically engineered organisms (in agriculture, in medicine).			
Hap	Evaluate the benefits and risks of using selective breeding and genetic engineering to produce new varieties and breeds.			

CB5 Health, Disease, and the Development of Medicines (Paper 1: Biology 1)

CB5a Health and disease

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Define the term health.			
5 th	Define the term disease.			
6 th	Describe how communicable and non-communicable diseases differ.			
7 th	Outline the role of the immune system in protecting against disease.			
8 th	Explain how disease can affect the immune system.			

CB5b Non-communicable disease

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Give examples of non-communicable diseases.			
4 th	Define the term malnutrition.			
5 th	Explain how diet can lead to malnutrition.			
6 th	Describe the link between alcohol and liver disease.			
7 th	Explain the effect of alcohol consumption on liver disease at local, national and global levels.			

CB5c Cardiovascular disease

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how obesity is measured (BMI and waist : hip calculations).			
6 th	Describe how obesity correlates with cardiovascular disease.			
6 th	Describe how smoking correlates with cardiovascular disease.			
6 th	Explain why exercise and diet affect obesity.			
8 th	Compare how cardiovascular diseases are treated			

CB5d Pathogens

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ch	Describe some problems and diseases caused by bacteria.			
5 th	Describe a disease caused by a virus.			
5 th	Describe a disease caused by a protist			
5 th	Describe a disease caused by a fungus.			
7 th	Explain how signs of a disease can be used to identify the pathogen.			

CB5e Spreading pathogens

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	State the ways in which pathogens can be spread.			
649	Give examples of pathogens that are spread in different ways (e.g. cholera bacteria by water, tuberculosis bacteria and chalara dieback fungi by air, malaria protist by vector, Helicobacter by mouth, Ebola by body fluids).			
7 th	Explain how the spread of different pathogens can be reduced or prevented.			

CB5f Physical and chemical barriers

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Explain how the spread of the STIs Chlamydia and HIV can be reduced or prevented.			
5 th	Give examples of physical barriers.			
5 th	Give examples of chemical barriers.			
6 th	Describe how physical barriers protect the body (e.g. skin, mucus and cilia).			
6 th	Describe how chemical barriers protect the body (e.g. lysozymes, hydrochloric acid).			

CB5g The immune system

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ^{ch}	State that the immune system protects the body by attacking pathogens.			
7th	Describe how antigens trigger the release of antibodies and the production of memory lymphocytes.			
7 th	Describe the role of antibodies in the immune response.			
7 th	Describe the role of memory lymphocytes in triggering a secondary response.			
8 th	Explain how immunisation protects against infection by a pathogen.			

CB5h Antibiotics

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Define the term antibiotic (as medicines that inhibit cell processes in bacteria).			
6 th	Explain why antibiotics are useful for treating bacterial infections (because they do not damage human cell processes).			
6 th	Explain why antibiotics cannot be used to treat infections by pathogens other than bacteria.			
6 th	Describe the stages of development of new medicines.			
7 th	Explain why each stage of the development of a new medicine is needed.			

CB6 Plant Structures and their Functions (Paper 4: Biology 2)

CB6a Photosynthesis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain why photosynthetic organisms are producers of biomass.			
6 th	Recall some substances produced from glucose and their roles in the plant.			
8 th	Summarise what happens in photosynthesis (including the use of a word equation).			
9th	Explain why photosynthesis is an endothermic reaction.			
6 th	Explain how a leaf and its cells are adapted for photosynthesis.			

CB6b Factors that affect photosynthesis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall what is meant by a rate of reaction.			
7 ^{ch}	Describe the effects of temperature, light intensity and carbon dioxide concentration on the rate of photosynthesis.			
9th	Explain the effects of limiting factors of photosynthesis.			
9th	Explain the effects of more than one factor on the rate of photosynthesis.			
9th	■ Describe how light intensity and rate of photosynthesis are related.			
94	Explain why the rate of photosynthesis is inversely proportional to the distance of a light source.			

CB6c Absorbing water and mineral ions

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how root hair cells are adapted to taking in water and mineral ions.			
6 th	Recall that substances can be transported by diffusion, osmosis and active transport.			
6 th	Describe what is meant by a concentration gradient.			
7 th	Explain why active transport is needed to transport some molecules.			
8 th	Explain how molecules move by osmosis.			

CB6d Transpiration and translocation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how xylem tissue is adapted to its functions.			
6 th	Explain how phloem tissue is adapted to its function.			
7 th	Describe how transpiration occurs.			
7 th	Describe how translocation occurs.			
911	Explain the effects of environmental factors on the rate of transpiration (light intensity, air movement, temperature, humidity).			
7 th	Describe how to measure the rate of transpiration.			

CB7 Animal Coordination, Control and Homeostasis (Paper 4: Biology 2)

CB7a Hormones

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	State where hormones are produced (in endocrine glands).			
6 th	Describe the general role of hormones in the body.			
6 th	Describe how hormones are transported around the body.			
6 th	Describe the production and release of some common hormones from their endocrine glands (pituitary gland, thyroid gland, pancreas, adrenal glands, ovaries and testes).			
6 th	Identify the target organs of some common hormones.			
7 th	Explain the importance of hormones.			

CB7b Hormonal control of metabolic rate

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe the effects of adrenalin on the body.			
7 th	Explain how adrenalin prepares the body for fight or flight.			
5 th	H Define metabolic rate.			
6 th	■ Describe the effect of thyroxine on metabolic rate.			
7 th	■ Describe how a negative feedback mechanism works.			
8 th	Explain how negative feedback controls the production of thyroxine.			
10 th	Explain why negative feedback mechanisms are important in living organisms.			

CB7c The menstrual cycle

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe what happens during the menstrual cycle.			
6 th	Describe the function of oestrogen in the menstrual cycle.			
6 th	Describe the function of progesterone in the menstrual cycle.			
7 th	Explain how barrier methods can be used as contraception.			
8 th	Explain how hormones can be used as contraception.			
9th	Compare, contrast and evaluate hormonal and barrier methods of contraception.			

CB7d Hormones and the menstrual cycle

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how changes in hormones affect the uterus wall, ovulation and menstruation.			
8 th	Explain how oestrogen, progesterone, FSH and LH interact in the menstrual cycle.			
6 th	Describe examples of Assisted Reproductive Technology (ART).			
8 th	Explain how clomifene is used to stimulate ovulation.			
8 th	Explain how hormones are used in IVF treatment.			

CB7e Control of blood glucose

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Define homeostasis.			
8 th	Explain why a constant internal environment is important.			
8 th	Explain the role of insulin in regulating blood glucose concentration.			
8 th	Explain the role of glucagon in regulating blood glucose concentration.			
7 th	Explain how type 1 diabetes is caused.			
7 th	Explain how type 1 diabetes can be controlled.			

CB7f Type 2 diabetes

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain how type 2 diabetes is caused.			
7 th	Explain how type 2 diabetes can be controlled.			
6 th	Describe the correlation between body mass and type 2 diabetes.			
7 th	Explain how BMI and waist : hip ratio are related to body mass.			
8 th	Evaluate the correlation between body mass and type 2 diabetes.			

CB8 Exchange and Transport in Animals (Paper 4: Biology 2)

CB8a Efficient transport and exchange

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall the names of substances that need to be transported into and out of the body.			
5 th	Describe the functions of the substances that are transported into the body.			
6 th	Describe the adaptations of the lungs for gas exchange.			
8 th	Calculate surface area : volume ratios.			
9 th	Explain the importance of surface area : volume ratios in transport systems.			

CB8b The circulatory system

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 ^{ch}	Recall the components and function of the circulatory system.			
5 th	Recall the functions of the different types of blood vessels.			
6 th	Describe the functions of the different types of blood cells (erythrocytes, phagocytes, lymphocytes).			
6 th	Describe the functions of blood platelets and plasma.			
7 th	Describe how the different blood vessels are adapted to their functions.			

CB8c The heart

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall the parts of the heart.			
6 th	Describe the flow of blood through the heart.			
7 ^{ch}	Explain how the heart is adapted for its function (valves, differing ventricle muscle thicknesses).			
7 ^{ch}	Recall and use the equation that relates cardiac output, stroke volume and heart rate.			

CB8d Cellular respiration

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain why organisms need to respire.			
6 th	Recall the word equation for aerobic respiration.			
6 th	Recall the word equation for anaerobic respiration in humans.			
8 th	Explain why respiration is an exothermic process.			
8 th	Compare aerobic and anaerobic respiration.			

CB9 Ecosystems and Material Cycles (Paper 4: Biology 2)

CB9a Ecosystems

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	State what is meant by the ecological terms community, population and habitat.			
4 th	Give examples of an ecosystem, a community, a population and a habitat.			
5 th	Describe the organisation of the components of an ecosystem (including populations, communities, habitats and abiotic factors).			
6 th	Describe how the interdependence of organisms in an ecosystem allows their survival.			
6 th	Explain how to estimate population size, including the use of quadrats.			

CB9b Abiotic factors and communities

Step	Learning outcome	Had a look	Nearly there	Nailed it!
3 rd	Give examples of abiotic factors.			
6th	Explain how communities are affected by abiotic factors (temperature, light, water, pollutants).			
7th	Explain how to investigate the effect of abiotic factors on the distribution of organisms using belt transects.			

CB9c Biotic factors and communities

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Give examples of biotic factors.			
6 th	Describe how competition can affect communities.			
6 th	Describe how predation can affect communities.			
7 th	Explain how predator–prey cycles affect communities.			
9th	Explain how the structure of a community can affect biodiversity.			

CB9d Parasitism and mutualism

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Define the term 'parasitism'.			
7 th	Define the term 'mutualism'.			
7 th	Describe how parasites are dependent on their hosts.			
7 th	Describe how hosts are harmed by parasites.			
7 th	Identify parasites and mutualists in examples.			
8 th	Explain how mutualists benefit from their relationship.			

CB9e Biodiversity and humans

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Define the term eutrophication.			
5 th	Describe examples of the introduction of non-indigenous species.			
6 th	Describe the advantages of fish farming.			
7 th	Explain how fish farming can affect ecosystems and biodiversity.			
7 th	Explain how the introduction of species can affect ecosystems and biodiversity.			
7 th	Explain how eutrophication can affect ecosystems and biodiversity.			

CB9f Preserving biodiversity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Define the term conservation.			
5 th	Explain what is meant by reforestation.			
5 th	Give examples of animal conservation.			
7 th	Explain how animal conservation can benefit biodiversity.			
7 th	Explain how reforestation can benefit biodiversity.			

CB9g The water cycle

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Give examples of materials that cycle through ecosystems.			
4 th	Describe the processes by which water cycles through abiotic parts of an ecosystem.			
5 th	Describe the processes by which water cycles through living organisms.			
5 th	Describe how drinking water is produced where water is plentiful.			
7 th	Explain how drinking water can be produced by desalination in areas of drought.			
7 th	Explain why water is important to living organisms.			

CB9h The carbon cycle

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Give examples of decomposers.			
6 th	Define the term decomposer.			
7 th	Describe the carbon cycle.			
7 th	Identify the key processes in the carbon cycle.			
8 th	Explain how carbon is cycled through the biotic and abiotic components of an ecosystem.			
8 th	Explain the importance of the carbon cycle (in balancing photosynthesis and respiration, and removal of wastes by decomposition).			

CB9i The nitrogen cycle

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe how plants use nitrates.			
7 th	Describe the different roles of bacteria in the nitrogen cycle.			
7 th	Explain how fertilisers increase the nitrate content of the soil.			
8 th	Explain why bacteria are important for soil fertility.			
8 th	Explain how crop rotation can increase the nitrogen content of the soil.			

CHEMISTRY

CC1 States of Matter (Paper 2: Chemistry 1)

CC1a States of matter

Step	Learning outcome	Had a look	Nearly there	Nailed it!
2 nd	Name the three states of matter, and the physical changes that occur between them.			
5 th	Describe the arrangements and movement of particles in the different states of matter.			
6 th	Use information to predict the state of a substance.			
5 th	Describe the relative energies of particles in the different states of matter.			
7 th	Explain why the movement and arrangement of particles change during changes of state.			
7 th	Explain why the energy of particles changes during changes of state.			

CC2 Methods of Separating and Purifying Substances (Paper 2: Chemistry 1)

CC2a Mixtures

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the differences between a pure substance and a mixture.			
5 th	Use melting point information to decide whether a substance is pure or is a mixture.			
6 th	Describe what happens to atoms at a pure substance's melting point.			
6 th	Interpret a heating curve to identify a melting point.			
7 th	Explain why the temperature does not change as a pure substance melts.			

CC2b Filtration and crystallisation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	State some mixtures that can be separated by filtration.			
4 th	State some mixtures that can be separated by crystallisation.			
6 th	Draw and interpret diagrams showing how filtration and crystallisation are done.			
6 th	Explain the formation of crystals during crystallisation.			
5 th	Explain how mixtures are separated by filtration.			
5 th	Explain ways of reducing risk when separating mixtures by filtration and crystallisation.			

CC2c Paper chromatography

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how some mixtures can be separated by chromatography.			
5 th	Identify pure substances and mixtures on chromatograms.			
5 th	Identify substances that are identical on chromatograms.			
6 th	Draw and interpret diagrams showing how chromatography is done.			
6 th	Explain how substances can be separated by chromatography.			
6 th	Calculate R _f values and use them to identify substances.			

CC2d Distillation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how to carry out, and explain what happens in, simple distillation.			
7 th	Distinguish between simple distillation and fractional distillation.			
7 th	Identify when fractional distillation should be used to separate a mixture.			
7 th	Describe how to carry out fractional distillation.			
9th	Explain how the products of fractional distillation are linked to the boiling points of the components.			
9th	Explain what precautions are needed to reduce risk in a distillation experiment.			

CC2e Drinking water

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Explain why water used in chemical analysis must not contain dissolved salts.			
5 th	Describe how fresh water can be produced from seawater.			
5 th	Describe the steps needed to make fresh water suitable for drinking.			
5 th	Suggest how to purify water when you know what it contains.			
8 th	Evaluate the hazards and control the risks present when purifying water.			

CC3 Atomic Structure (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC3a Structure of an atom

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Describe how Dalton's ideas about atoms have changed.			
8 th	Describe how the subatomic particles are arranged in an atom.			
8 th	Explain how atoms of different elements are different.			
7 th	Recall the charges and relative masses of the three subatomic particles.			
8 th	Explain why all atoms have no overall charge.			
8 th	Describe how the size of an atom compares to the size of its nucleus.			

CC3b Atomic mass and number

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State where most of the mass of an atom is found.			
7 th	State the meaning of atomic number.			
7 th	State the meaning of mass number.			
8 th	Describe how the atoms of different elements vary.			
8 th	State the number of electrons in an atom from its atomic number.			
8 th	Calculate the numbers of protons, neutrons and electrons using atomic and mass numbers.			

CC3c Atomic mass and number

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State what is meant by an isotope.			
7 th	Identify isotopes from information about the structure of atoms.			
8 th	Calculate the numbers of protons, neutrons and electrons using atomic numbers and mass numbers.			
9 th	Explain why the relative atomic mass of many elements is not a whole number.			
100	■ Calculate the relative atomic mass of an element from the relative masses and abundances of its isotopes.			

CC4 The Periodic Table (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC4a Elements and the periodic table

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the chemical symbols of some common elements.			
6 th	Describe how Mendeleev arranged elements into a periodic table.			
7 th	Describe how Mendeleev predicted the existence and properties of some elements yet to be discovered.			
8 th	Explain how Mendeleev's early ideas were supported by later evidence.			

CC4b Atomic number and the periodic table

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain some problems Mendeleev had when ordering the elements.			
6 th	Explain the meaning of the term 'atomic number'.			
6 th	Describe how the elements are arranged in the modern periodic table.			
6 th	Recall the positions of metals and non-metals in the periodic table.			

CC4c Electronic configurations and the periodic table

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State what the term 'electronic configuration' means.			
8 th	Show electronic configurations in the form 2.8.1 and as diagrams.			
9 th	Predict the electronic configurations of the elements hydrogen to calcium.			
9 th	Explain the links between an element's position in the periodic table and its electronic configuration.			

CC5 Ionic Bonding (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC5a Ionic bonds

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the formulae of simple ions.			
8 th	Explain how cations and anions are formed.			
8 th	Use dot and cross diagrams to explain how ionic bonds are formed.			
8 th	Explain the difference between an atom and an ion.			
9th	Calculate the numbers of protons, neutrons and electrons in simple ions.			
9th	Explain the formation of ions in groups 1, 2, 6 and 7 of the periodic table.			

CC5b Ionic lattices

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the formulae of common polyatomic ions, and the charges on them.			
7 th	Interpret the use of –ide and –ate endings in the names of compounds.			
7 th	Name ionic compounds using –ide and –ate endings.			
8 th	Work out the formula of an ionic compound from the formulae of its ions.			
8 th	Describe the structure of ionic compounds.			
8 th	Explain how ionic compounds are held together.			

CC5c Properties of ionic compounds

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the properties of ionic compounds.			
7 th	Explain why ionic compounds have high melting points and high boiling points.			
7 th	Explain why ionic compounds conduct electricity when they are molten and in aqueous solution.			
7 th	Explain why ionic compounds do not conduct electricity as solids.			
7 th	Identify ionic compounds from data about their properties.			

CC6 Covalent bonding (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC6a Covalent bonds

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain how covalent bonds are formed.			
5 th	Recall the names of some common molecular elements.			
5 th	Recall the names of some common molecular compounds.			
6 th	State the bonding that is found in molecules.			
6 th	State the approximate size (order or magnitude) of atoms and small molecules.			
8 th	Explain the formation of covalent bonds using dot and cross diagrams.			

CC7 Types of Substance (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC7a Molecular compounds

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall examples of common covalent, simple molecular compounds.			
6 th	Describe the general properties of covalent, simple molecular compounds.			
8 th	Explain why covalent, simple molecular compounds have low melting and boiling points.			
9th	Explain why covalent, simple molecular compounds are poor conductors of electricity.			
7 th	Describe the structure of a polymer.			

CC7b Allotropes of carbon

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ch	Recall some allotropes of carbon.			
7 th	Describe the basic differences between covalent, simple molecules and giant covalent structures.			
7 th	Describe the structures of diamond, graphite, fullerenes and graphene.			
6 th	Describe the properties of diamond, graphite, fullerenes and graphene.			
9th	Explain the properties and uses of diamond and graphite in terms of their structure and bonding.			
9th	Explain the properties of fullerenes and graphene in terms of their structure and bonding.			

CC7c Properties of metals

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe the particles and how they are arranged in metals.			
7 th	Explain why metals are malleable.			
7 th	Explain why metals conduct electricity.			
3 rd	Describe the typical properties of metals.			
3 rd	Describe the typical properties of non-metals.			

CC7d Bonding metals

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Give examples of ionic; covalent, simple molecular; covalent, giant molecular; and metallic substances.			
7 th	Describe how the different types of bonds and structures are formed.			
8**	Explain how the structure and bonding of a substance is linked to its physical properties. (Relative melting point and boiling point, relative solubility in water and ability to conduct electricity, as solids and in solution.)			
8 th	Explain why we use models to represent structure and bonding.			
8 ^{ch}	Represent structures and bonding using a variety of different models (dot and cross, ball and stick, 2D, 3D).			
9 th	Describe the limitations of the different models used to represent structure and bonding (dot and cross, ball and stick, 2D, 3D).			

CC8 Acids and Alkalis (Paper 2: Chemistry 1)

CC8a Acids, alkalis and indicators

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Describe what the main hazard symbols mean.			
5 th	Describe the safety precautions that should be observed when handling different acids and alkalis.			
4 th	Name the ions present in all acidic and all alkaline solutions.			
5 th	State the pH values associated with acidic, alkaline and neutral solutions.			
5 th	Describe the effect of acids and alkalis on common indicators.			
5 th	Explain the link between pH and the concentration of ions in acids and alkalis.			

CC8b Looking at acids

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ch	■ Describe the relationship between hydrogen ion concentration and pH.			
5 ^{ch}	Explain the difference between a dilute and concentrated solution (in terms of the amount of solute present).			
7 th	Explain the difference between strong and weak acids (in terms of the degree of dissociation of the acid molecules).			
7 th	Explain how the pH and reactivity of an acid depend on the concentration and the strength of the acid.			

CC8c Bases and salts

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how a base reacts in a neutralisation reaction.			
6 th	Describe what happens when an acid reacts with a metal oxide.			
7 th	Write word equations for the reactions of acids and metal oxides.			
8 th	Write symbol equations for the reactions of acids and metal oxides.			
6 th	Explain what happens during a neutralisation reaction.			
6 th	Describe the steps involved in preparing a soluble salt from an acid and an insoluble reactant.			
6 th	Explain why: an excess of insoluble reactant is used when preparing a soluble salt the excess reactant is removed when			
	preparing a soluble salt the remaining solution contains only a salt and water, when preparing a soluble salt from an acid and an insoluble reactant.			

CC8d Alkalis and balancing equations

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the chemical formulae of some common compounds.			
6 th	Recall and use state symbols.			
9 th	Balance chemical equations.			
4 th	Recall that alkalis are soluble bases.			
6 th	Describe the reactions of alkalis with acids.			

CC8e Alkalis and neutralisation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain what happens to the ions from acids and alkalis during neutralisation.			
6 th	Explain why titration is used to prepare soluble salts.			
6 th	Describe how to carry out an acid–alkali titration.			

CC8f Reactions of acids with metals and carbonates

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9 th	₩ Write balanced ionic equations.			
7 th	Explain the general reaction between an acid and a metal to produce a salt and hydrogen.			
7 th	Explain the general reaction between an acid and a metal carbonate to produce a salt, water and carbon dioxide.			
5 th	Describe the test for hydrogen.			
5 th	Describe the test for carbon dioxide.			

CC8g Solubility

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall the general rules for the solubility of common substances in water.			
6 th	Predict whether or not a precipitate will form from two solutions.			
6 th	Name the precipitate formed in a reaction.			
6 th	Describe how to prepare a pure, dry sample of an insoluble salt.			

CC9 Calculations involving masses (Paper 2: Chemistry 1 and Paper 5: Chemistry 2)

CC9a Masses and empirical formulae

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Calculate the relative formula mass of a substance from relative atomic masses.			
8 th	Calculate the empirical formula of a compound from the masses of the elements it contains.			
8 th	Explain the difference between an empirical formula and a molecular formula.			
7 th	Deduce the empirical formula from a molecular formula.			
8 th	Deduce the molecular formula for a compound from its empirical formula and its relative formula mass.			
7 th	Describe an experiment to determine the empirical formula for a compound.			

CC9b Conservation of mass

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain the law of conservation of mass in a closed system.			
6 th	Explain the law of conservation of mass in a non-enclosed system.			
8 th	Calculate the mass of product formed from a given mass of reactant, using a balanced equation.			
8 th	Calculate the mass of a reactant needed to produce a given amount of product, using a balanced equation.			
6 th	Calculate the concentration of a solution in g dm ⁻³ .			

CC9c Moles H

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	H Describe what is meant by a mole of particles.			
8 th	☐ Calculate the number of moles of particles in a given mass of a certain substance and vice versa.			
9 th	☐ Calculate the number of particles in a given number of moles or mass of a substance and vice versa.			
7 th	Explain that the mass of a product formed in a reaction is controlled by the mass of reactant that is not in excess.			
9 th	Deduce the balanced equation for a reaction from the masses of reactants and/or products.			

CC10 Electrolytic Processes (Paper 2: Chemistry 1)

CC10a Electrolysis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	State the meaning of the term 'electrolyte'.			
7 th	Outline what happens during electrolysis.			
7 th	Explain the movement of the ions during electrolysis.			
8 th	■ Write half equations for the reactions at the electrodes.			
9 th	Explain the meaning of oxidation and reduction in terms of the movement of electrons.			
8 th	State the electrodes at which oxidation and reduction occur.			

CC10b Products from electrolysis

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the products formed from the electrolysis of a variety of common compounds and solutions (copper chloride solution, sodium chloride solution, sodium sulfate solution, acidified water, molten lead bromide).			
8.0	Explain the formation of the products in the electrolysis of a variety of common compounds and solutions (copper chloride solution, sodium chloride solution, sodium sulfate solution, acidified water, molten lead bromide).			
8th	Predict the products formed from the electrolysis of a molten, binary, ionic compound.			
8 th	Explain how the electrolysis of copper sulfate solution using copper electrodes can be used to purify copper.			

CC11 Obtaining and Using Metals (Paper 2: Chemistry 1)

CC11a Reactivity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the reactions of common metals with water and acids.			
5 th	Describe the reactions of metals with salt solutions.			
8 th	Explain why displacement reactions are redox reactions.			
8 th	Deduce the order of metals in the reactivity series from their reactions with water, acids and salt solutions.			
9 th	Explain the reactivity series in terms of the tendency of different metal atoms to form cations.			

CC11b Ores

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall the meaning of the term 'ore'.			
4 th	Recall some metals that are found uncombined in the Earth's crust.			
7 th	Explain how and why some metals are extracted from their ores by heating with carbon.			
8 th	Explain how and why some metals are extracted from their ores by electrolysis.			
7 th	■ Describe two biological methods of metal extraction.			
10 ^{ch}	Evaluate biological methods of metal extraction.			

CC11c Oxidation and reduction

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9**	Explain why reactions occurring at the electrodes during electrolysis are redox reactions.			
8 th	Describe the meanings of oxidation and reduction in terms of oxygen.			
9 th	Explain which substance has been oxidised and which substance has been reduced in a reaction.			
7 th	Recall that all metals are extracted by reduction of their ores.			
8 th	Explain how the position of a metal in the reactivity series is related to its resistance to oxidation.			

CC11d Life cycle assessment and recycling

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	State the advantages and disadvantages of recycling a metal.			
5 th	Describe a process where a material or product is recycled for a different use.			
8 th	Evaluate the advantages and disadvantages of recycling a material or product to decide whether recycling is a viable option.			
5 ^{ch}	Describe the four stages in carrying out a life cycle assessment (LCA) of a material or product.			
8 th	Evaluate data from a life cycle assessment of a material or product.			

CC12 Reversible Reactions and Equilibria (Paper 2: Chemistry 1)

CC12a Dynamic equilibrium

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ch	Describe what happens in reversible reactions.			
7 th	Explain the use of the symbol ≠ in chemical equations.			
7 th	Explain what is meant by dynamic equilibrium.			
7 th	Describe the formation of ammonia.			
9th	State the conditions used for the Haber process.			
9 th	Describe how changing the temperature, pressure and concentration all affect the relative amount of substances in an equilibrium mixture.			

CC13 Groups in the Periodic Table (Paper 5: Chemistry 2)

CC13a Group 1

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain the classification of alkali metals, halogens and noble gases, into groups in the periodic table.			
6 th	Describe the main physical properties of alkali metals.			
6 th	Describe the reactions of lithium, sodium and potassium with water.			
9th	Write word, balanced and H ionic equations (including state symbols) for the reactions of alkali metals.			
9 th	Describe the pattern of reactivity of the alkali metals.			
10	Explain how the electronic configurations of the atoms of alkali metals affect their reactivity.			

CC13b Group 7

Step	Learning outcome	Had a look	Nearly there	Nailed it!
3 rd	Recall the appearance of chlorine, bromine and iodine at room temperature.			
8 th	Describe the trends in colour, melting point and boiling point of chlorine, bromine and iodine down the group, and use these to predict physical properties of other halogens.			
5 th	Describe the chemical test for chlorine gas.			
8 th	Describe the trends in the reactions of halogens with metals, and use this to predict reactions of other halogens.			
9 th	Write word and balanced chemical equations, including state symbols, for the reactions of halogens with metals.			
6 th	Describe hydrogen halides and their chemical properties.			

CC13c Halogen reactivity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe the relative reactivity of halogens.			
8 th	Explain how the reactivity of halogens can be worked out from displacement reactions.			
9 th	Write balanced chemical equations, including state symbols, for the displacement reactions of halogens.			
10-гр	Explain how displacement reactions are examples of redox reactions.			
8 th	Write ionic equations, including state symbols, for displacement reactions of halogens.			
10-гр	Explain the order of reactivity of halogens (using electronic configurations).			

CC13d Group 0

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Explain why noble gases are chemically inert by referring to their electronic configuration.			
3rd	Describe uses of noble gases linked with their properties.			
6 th	Describe the trends in the physical properties of the noble gases.			
8 th	Use trends in physical properties to predict the physical properties of other noble gases.			

CC14 Rates of Reaction (Paper 5: Chemistry 2)

CC14a Rates of reaction

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe different changes that can occur as a reaction proceeds.			
7 th	Suggest different experimental methods to investigate rates of reaction (e.g. measurements of mass of reactants against time, volume of gas released against time, concentration of reactant or product against time).			
7 th	Use graphs of changes (in mass, volume or concentration of reactant or product) against time, to interpret what is happening during reactions.			

CC14b Factors affecting reaction rates

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Explain what has to happen for reactions to take place.			
9th	Explain why changes in the energy of particles affect rates of reaction.			
9 th	Explain why changes in the frequency of collisions between particles affect the rate of reaction.			
9 th	Explain why changes in temperature, concentration, surface area and pressure affect the rate of reaction (surface area for solids, pressure for gases only).			
8 th	Describe ways of speeding up or slowing down chemical reactions.			

CC14c Catalysts and activation energy

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe what a catalyst does.			
7 th	Explain how catalysts are useful.			
8 th	Explain what the activation energy of a reaction is.			
9 th	Explain how catalysts speed up chemical reactions.			
7 th	Describe what enzymes are.			
6 th	Name one or more examples of enzymes.			

CC15 Heat energy changes in chemical reactions (Paper 5: Chemistry 2)

CC15a Exothermic and endothermic reactions

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall some examples of exothermic and endothermic changes.			
7 th	Describe how heat changes in solution may be determined.			
8 th	Describe the differences between exothermic and endothermic changes.			

CC15b Energy changes in reactions

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9 th	Describe exothermic and endothermic reactions in terms of energy changes when bonds are broken and formed.			
II th	■ Use bond energies to calculate energy changes in reactions.			
9 th	Explain the meaning of activation energy.			
10 th	Draw and label reaction profiles.			

CC16 Fuels (Paper 5: Chemistry 2)

CC16a Hydrocarbons in crude oil and natural gas

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Recall the meaning of the term hydrocarbon.			
6 th	Describe the compounds found in crude oil.			
5 th	Describe the importance of crude oil for the petrochemical industry.			
6 th	Explain why crude oil is a finite resource.			
4 th	Recall the names of some common fossil fuels.			

CC16b Fractional distillation of crude oil

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how crude oil is separated by fractional distillation.			
8 th	Explain how fractional distillation of crude oil works.			
6 th	Recall the names and uses of fractions from crude oil.			
8 th	Describe how fractions differ from each other.			
8 th	Explain why the properties of different fractions differ.			

CC16c The alkane homologous series

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe that oil fractions mostly contain alkanes.			
7 th	Describe the main features of an homologous series.			
8 th	Explain why alkanes form an homologous series.			

CC16d Complete and incomplete combustion

Step	Learning outcome	Had a look	Nearly there	Nailed it!
- O	Describe the complete combustion of hydrocarbon fuels.			
7 th	Explain the production of harmful products during the incomplete combustion of hydrocarbon fuels.			
7 th	Explain why carbon monoxide is toxic.			
/cii	Describe the problems caused by incomplete combustion.			

CC16e Combustible fuels and pollution

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how some hydrocarbon fuels produce sulfur dioxide in use.			
5 th	Recall the names of the pollutants responsible for acid rain.			
5 th	Describe some effects of acid rain.			
6 th	Explain why oxides of nitrogen are produced when fuels are burned in engines.			

CC16f Breaking down hydrocarbons

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9th	Evaluate hydrogen as an alternative fuel to petrol for cars.			
6 th	Describe what happens during cracking.			
_ g	Explain why alkanes are saturated and alkenes are unsaturated.			
7th	Explain why cracking is necessary.			

CC17 Earth and Atmospheric Science (Paper 5: Chemistry 2)

CC17a The early atmosphere

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how the Earth's early atmosphere was formed.			
5 th	State the names and relative amounts of the gases found in the Earth's early atmosphere.			
7 th	Draw conclusions from evidence about the Earth's early atmosphere.			
6 th	Explain how the oceans are thought to have formed.			

CC17b The changing atmosphere

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how the formation of the oceans influenced the composition of the atmosphere.			
7 th	Explain how photosynthetic organisms (including plants) changed the composition of the atmosphere.			
5 th	State the chemical test for oxygen.			

CC17c The atmosphere today

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall the names of significant greenhouse gases.			
7 th	Describe the processes involved in the greenhouse effect.			
6 th	Describe how human activity increases the concentration of greenhouse gases.			
1000	Evaluate the correlation between atmospheric carbon dioxide concentrations and fossil fuel use.			
100	Evaluate the evidence for increased atmospheric greenhouse gas concentrations being part of the cause of global warming and climate change.			

CC17d Climate change

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Suggest possible effects on the climate of increased levels of carbon dioxide and methane.			
6 th	Describe how human activity leads to increased carbon dioxide levels.			
6 th	Describe how human activity leads to increased methane levels.			
6 th	Describe the projected effects of climate change.			
7 th	Describe how the potential harmful effects of climate change can be addressed and limited.			

PHYSICS

CP1 Motion (Paper 3: Physics 1)

CP1a Vectors and scalars

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Describe the difference between weight and mass.			
7 th	Explain the difference between a vector and a scalar quantity.			
7 th	Describe the difference between displacement and distance.			
7 th	Describe the difference between velocity and speed.			
6 th	Define the terms: acceleration, force, momentum, energy.			

CP1b Distance/time graphs

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall formulae relating distance, speed and time.			
7 th	Use formulae relating distance, speed and time.			
5 th	Recall typical speeds for walking, running, cycling and travelling by car.			
6 th	Interpret distance/time graphs (including recognising what the steepness of the line tells you).			
7 th	Represent journeys on distance/time graphs.			
8 th	Determine speed from the gradient of a distance/time graph.			

CP1c Acceleration

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the formula relating acceleration, velocity and time.			
8 th	Use the formula relating acceleration, velocity and time.			
6 th	Recall the formula relating acceleration, velocity and distance.			
8 th	Use the formula relating acceleration, velocity and distance.			
6 th	Recall the acceleration in free fall.			
8 th	Estimate the magnitudes of some everyday accelerations.			

CP1d Velocity/time graphs

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Represent journeys on velocity/time graphs.			
7 th	Interpret velocity/time graphs qualitatively.			
8 th	Calculate uniform accelerations from the gradients of velocity/time graphs.			
9 th	Determine the distance travelled from the area under a velocity/time graph.			

CP2 Motion and Forces (Paper 3: Physics 1)

CP2a Resultant forces

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain the difference between scalar and vector quantities.			
7 th	Use arrows to represent the direction and magnitude of forces.			
6 th	Define a resultant force.			
6 th	Calculate resultant forces.			
6 th	Explain whether forces on an object are balanced or unbalanced.			

CP2b Newton's First Law

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ^{ch}	Describe the effect of balanced forces on moving and stationary objects.			
6 th	Describe the effect of a non-zero resultant force on moving and stationary objects.			
7 th	Describe circular motion at constant speed as a changing velocity and hence as an acceleration.			
7 th	Describe the force needed to keep an object moving in a circular path.			
8 th	Give some examples of objects moving in circular paths and the type of centripetal force involved.			

CP2c Mass and weight

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Describe the difference between mass and weight.			
4 th	List the factors that determine the weight of an object.			
4 th	Recall the formula for calculating weight.			
6 th	Calculate weights using the formula.			
7 th	Change the subject of the weight formula to calculate mass or gravitational field strength.			

CP2d Newton's Second Law

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe what an acceleration is.			
6 th	List the factors that affect the acceleration of an object.			
6 th	Recall the formula that relates the factors affecting acceleration.			
8 th	Use the formula relating force, mass and acceleration.			
9th	Change the subject of the formula relating force, mass and acceleration.			
7 th	Explain what inertial mass means.			

CP2e Newton's Third Law

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe what Newton's Third Law says.			
7 th	Recall the meaning of 'equilibrium situation'.			
8 th	Identify action–reaction pairs in familiar situations.			
8 th	Distinguish between action–reaction pairs and balanced forces.			
8 th	Describe how objects affect each other when they collide.			

CP2f Momentum

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe the factors that affect the momentum of an object.			
9 th	Calculate the momentum of moving objects.			
7 th	Recall what happens to momentum during a collision.			
9 th	Use the idea of conservation of momentum to calculate velocities of objects after collisions.			
10 ^{ch}	Calculate the force needed to produce a change in momentum in a given time.			

CP2g Stopping distances

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe how human reaction times are measured.			
5 th	Recall typical human reaction times and the factors that affect them.			
5 th	Describe the link between stopping distance, thinking distance and braking distance.			
5 ^{ch}	Recall the factors that affect stopping distances.			
6 th	Describe how different factors affect stopping distances.			

CP2h Crash hazards

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Explain the meaning of a 'large deceleration'.			
6 th	Describe the dangers caused by large decelerations.			
7 th	Explain why large decelerations cause dangers.			
7 th	Recall some typical forces involved in road collisions.			
9 th	■ Use knowledge of changes in momentum to estimate the forces involved in road collisions.			

CP3 Conservation of Energy (Paper 3: Physics 1)

CP3a Energy stores and transfers

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain, using examples, that energy is conserved.			
5 ^{ch}	Give examples of energy being moved between different stores.			
6 th	Interpret diagrams that represent energy transfers.			
7 th	Represent energy transfers using diagrams.			
7 th	Describe what happens to wasted energy in energy transfers.			

CP3b Energy efficiency

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Explain some ways in which energy is transferred wastefully by mechanical processes.			
7 th	Explain some ways of reducing unwanted energy transfers in mechanical processes.			
6 th	Define what efficiency means.			
7 th	Explain how efficiency can be increased.			
9th	Recall and use the formula for calculating energy efficiency.			

CP3c Supplying electricity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe what is meant by electrical resistance.			
8 th	Explain how energy can be wasted in electrical appliances.			
6 th	Describe how the National Grid transmits electricity around the country.			
7 th	Explain why step-up and step-down transformers are used in the National Grid.			
7 th	Explain how wasteful energy transfers can be reduced in electrical appliances.			

CP3d Keeping warm

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the ways in which energy can be transferred by heating.			
7 ^{ch}	Describe ways of reducing unwanted energy transfers using thermal insulation.			
7 ^{ch}	Explain how different ways of reducing energy transfer by heating work.			
5 th	Define the meaning of thermal conductivity.			
6 th	Describe the effects of the thickness and thermal conductivity of the walls of a building on its rate of cooling.			

CP3e Non-renewable resources

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	List the non-renewable energy resources in use today.			
5 th	Describe the advantages and disadvantages of non-renewable energy resources.			
7 th	Compare the advantages and disadvantages of non-renewable energy resources.			
6 th	Explain how the use of non-renewable energy resources is changing.			

CP3f Renewable resources

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	List the renewable energy resources in use today.			
5 th	Describe the source of energy for different renewable resources.			
5 ^{ch}	Describe the ways in which the different energy resources are used.			
7 th	Explain why we cannot use only renewable energy resources.			
6 th	Explain how the use of renewable energy resources is changing.			

CP4 Waves (Paper 3: Physics 1)

CP4a Describing waves

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 ch	Recall that waves transfer energy and information but do not transfer matter.			
5 th	Describe waves using the terms frequency, wavelength, amplitude, period and velocity.			
6 th	Describe the differences between longitudinal and transverse waves.			
4 th	Give examples of transverse and longitudinal waves.			

CP4b Waves velocities

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the equation relating wave speed, frequency and wavelength			
8 th	Use the equation relating wave speed, frequency and wavelength.			
6 th	Recall the equation relating wave speed, distance and time.			
8 th	Use the equation relating wave speed, distance and time.			
7 th	Describe how to measure the velocity of sound in air.			
7 th	Describe how to measure the velocity of waves on the surface of water.			

CP4c Refraction

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe what refraction is.			
5 th	Describe how the direction of a wave changes when it goes from one material to another.			
6 th	Explain some effects of the refraction of light (explanations in terms of changing speeds are not expected).			
7 th	Explain how a change in wave speed can cause a change in direction.			

CP5 Light and the Electromagnetic Spectrum (Paper 3: Physics 1)

CP5a Electromagnetic waves

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall examples of electromagnetic waves.			
5 th	Describe the common features of electromagnetic waves.			
5 th	Describe the transfer of energy by electromagnetic waves.			
5 th	Describe the range of electromagnetic waves that our eyes can detect.			
7 th	Describe an effect caused by the different velocities of electromagnetic waves in different substances.			

CP5b The electromagnetic spectrum

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall the groups of waves in the electromagnetic spectrum in order.			
5 th	Recall the colours of the visible spectrum in order.			
5 th	Describe how the waves in the electromagnetic spectrum are grouped.			
7 th	Describe some differences in the ways that different parts of the electromagnetic spectrum are absorbed and transmitted.			
8 th	Describe some differences in the ways that different parts of the electromagnetic spectrum are refracted and reflected.			

CP5c Using the long wavelengths

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	■ Describe how long wavelength electromagnetic waves are affected by different substances.			
7 th	Explain the effects caused by long wavelength electromagnetic waves travelling at different velocities in different substances.			
6 th	Describe some uses of radio waves.			
6 th	Describe some uses of microwaves.			
6 th	Describe some uses of infrared.			
6 th	Describe some uses of visible light.			
6 th	Describe how radio waves are produced and detected by electrical circuits.			

CP5d Using the short wavelengths

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how short wavelength electromagnetic waves are affected by different substances.			
7 th	Explain the effects caused by short wavelength electromagnetic waves travelling at different velocities in different substances.			
6 th	Describe some uses of ultraviolet radiation.			
6 th	Describe some uses of X-rays.			
6 th	Describe some uses of gamma rays.			

CP5e EM radiation dangers

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how the potential danger of electromagnetic radiation depends on its frequency.			
6 th	Describe the harmful effects of microwave and infrared radiation.			
6 th	Describe the harmful effects of ultraviolet radiation, X-rays and gamma rays.			
7 th	Recall the nature of radiation produced by changes in atoms and their nuclei.			
7 th	Recall that absorption of radiation can cause changes in atoms and their nuclei.			

CP6 Radioactivity (Paper 3: Physics 1)

CP6a Atomic models

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe the structure of an atom (in terms of nucleus and electrons).			
7 th	State where most of the mass of an atom is found.			
7 th	State the sizes of atoms and small molecules.			
8 th	Describe an early model of the atom.			
8 th	Describe how and why our model of the atom has changed over time, including the plum pudding model and the Rutherford alpha particle scattering.			

CP6b Inside atoms

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	State what is meant by an isotope.			
8 th	Represent isotopes using symbols.			
8 th	Explain how atoms of different elements are different (in terms of numbers of electrons and protons).			
7 th	Recall the charges and relative masses of the three subatomic particles.			
8 th	Explain why all atoms have no overall charge.			

CP6c Electrons and orbits

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe where electrons are found inside atoms (in terms of shells).			
8 th	Describe when electrons can change orbit.			
7 th	Recall what an ion is.			
8 th	Describe how ionisation occurs.			
8 th	Describe some of the evidence for the Bohr model of the atom.			

CP6d Background radiation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9th	Explain what background radiation is.			
9 th	Describe how radiation measurements need to be corrected for background radiation.			
8 th	List some sources of background radiation.			
8 th	Describe how photographic film can be used to detect radioactivity.			
9th	Describe how a Geiger-Müller tube works.			
849	Describe how the amount of radioactivity can be measured (in terms of the darkness of photographic film or by attaching a counter to a GM tube).			

CP6e Types of radiation

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	List five types of radiation that are emitted in random processes from unstable nuclei.			
8 th	State that the five types of radiation are ionising radiations.			
8 th	Describe what alpha and beta particles are.			
8 th	Describe the nature of gamma radiation.			
110	Compare the penetrating abilities of alpha, beta and gamma radiation.			
H	Compare the ionisation abilities of alpha, beta and gamma radiation.			

CP6f Radioactive decay

Step	Learning outcome	Had a look	Nearly there	Nailed it!
9 th	Describe the process of β^- decay.			
9 th	Describe the process of β^+ decay.			
10 th	Explain how the proton and mass numbers are affected by different kinds of radioactive decay.			
9th	Describe what happens during nuclear rearrangement after radioactive decay.			
10 th	Balance nuclear equations for mass and charge.			

CP6g Half-life

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Describe how the activity of a substance changes over time.			
8 th	State how half-life can be used to describe the changing activity of a substance.			
8 th	Recall the unit of activity.			
8 th	Describe how half-life can be used to work out how much of a substance will decay in a certain time.			
10 th	Carry out calculations involving half-life.			

CP6h Dangers of radioactivity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Describe the hazards of ionising radiation in terms of tissue damage and possible mutations.			
9 th	Explain the precautions taken to reduce the risks from radiation and ensure the safety of patients exposed to radiation.			
9 th	Explain the precautions taken to reduce the risks from radiation and protect people who work with radiation.			
9 th	Describe the differences between contamination and irradiation effects.			
II th	Compare the hazards of contamination and irradiation.			

CP7 Energy – Forces doing work (Paper 6: Physics 2)

CP7a Work and power

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe some ways in which the energy of a system can be changed.			
6 th	Measure the work done by a force.			
8 th	Recall and use the equation linking work done, force and distance.			
6 th	Explain what power means.			
8 th	Recall and use the equation linking power, work done and time.			

CP8 Forces and their Effects (Paper 6: Physics 2)

CP8a Objects affecting each other

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the effect of a gravitational field on objects.			
5 th	Describe the effects of magnetic fields on objects.			
5 th	Describe the forces that can occur when objects are in contact with each other.			
6 th	Describe the effects of electrostatic fields on objects.			
8 th	Describe how pairs of forces occur when objects affect each other.			
8 th	Use examples to explain the difference between vector and scalar quantities.			

CP8b Vector diagrams

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe how to resolve forces.			
8 th	Use scale drawings to work out the net force on an object.			
8 th	Draw free body diagrams to represent the forces on an object.			
9 th	Explain what happens in situations where several forces are acting on an object.			

CP9 Electricity and Circuits (Paper 6: Physics 2)

CP9a Electric circuits

Step	Learning outcome	Had a look	Nearly there	Nailed it!
7 th	Describe the basic structure of an atom (positions, relative masses and relative charges of protons, neutrons and electrons).			
3rd	Recognise the circuit symbols for a range of common electrical components (cells, including batteries, switches, voltmeters, ammeters and lamps).			
4 ^{ch}	Draw diagrams for circuits containing common electrical components, using conventions for positive and negative terminals.			
5 th	Describe and explain the difference between the brightness of identical lamps in series and parallel circuits.			
5 th	Describe and explain the effects of different numbers of identical lamps, cells and switches in series and parallel circuits.			

CP9b Current and potential difference

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Describe how to measure voltage.			
5 th	Define the term 'potential difference'.			
4 th	Describe how to measure current.			
4 ^{ch}	Describe the conditions needed to produce an electric current. (A complete circuit and a source of voltage/potential difference.)			
4 th	Describe the behaviour of current at a junction.			

CP9c Current, charge and energy

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain the link between the potential difference (voltage) across a battery or a component, the charge passing through it and the amount of energy transferred.			
6 th	Recall that the unit of potential difference is the volt and explain it in terms of units of energy and charge (a potential difference of one joule per coulomb).			
8 ^{ch}	Recall and use the equation to calculate the energy transferred, the charge that flows or the potential difference. ($E = Q \times V$)			
5 th	Explain the link between electric current and electric charge.			
5 th	Explain electric current in metals in terms of electrons.			
7 th	Recall and use the equation to calculate the charge that flows, the current or the time the current flows. $(Q = I \times t)$			

CP9d Resistance

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain the link between resistance and current in a circuit.			
5 th	Define the resistance of a component or circuit $(R = V/I)$.			
8 th	Recall and use the equation to calculate the potential difference, the current or the resistance $(V = I \times R)$.			
6 th	Explain the difference in resistance when two resistors are connected in series or in parallel.			
7 th	Calculate the currents, potential differences and resistances in series circuits.			
5 th	Explain the design and construction of series circuits for testing and measuring.			

CP9e More about resistance

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how current changes with potential difference in fixed resistors.			
6 th	Explain how current and resistance change with potential difference in filament lamps.			
6 th	Explain how current and resistance change with potential difference in diodes, including light-emitting diodes (LEDs).			
6 th	Describe how the resistance of a light- dependent resistor (LDR) varies with changing light intensity.			
6 th	Describe how the resistance of a thermistor varies with changing temperature. (negative temperature coefficient only)			
5 th	Describe the uses of diodes, LDRs and thermistors.			

CP9f Transferring energy

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Describe the energy transfer that occurs when a current passes through a resistor.			
7 th	Use the electron and ion model and the idea of electrical work to explain the energy transfer in a resistor and the resulting dissipation of energy in the surroundings.			
5 th	Explain how unwanted energy transfers in wires can be avoided.			
5 th	Recall the advantages of the heating effect of an electric current.			
5 th	Recall the disadvantages of the heating effect of an electric current.			
8 th	Use the equation $E = I \times V \times t$ to calculate the energy transferred, the current, the potential difference or the time.			

CP9g Power

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Define power and the units used to measure it. (energy transferred per second in watts)			
8 th	Recall and use the equation to calculate the power, the energy transferred or the time taken. $(P = E/t)$			
6 th	Explain how power transfer depends on the potential difference across a device and the current through it.			
8 th	Recall and use the equation to calculate the electrical power, the current or the potential difference. $(P = I \times V)$			
8 th	Recall and use the equation to calculate the electrical power, the current or the resistance. $(P = P \times R)$			

CP9h Transferring energy by electricity

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe energy transfers from d.c. batteries and the a.c. mains supply to motors and heaters.			
6 th	Explain the difference between direct and alternating voltage.			
6 th	Compare alternating and direct current (in terms of movement of charge).			
5 th	Recall the frequency and voltage of the UK domestic supply.			
6 th	Describe the power ratings of some domestic electrical appliances and changes in stored energy when they are in use.			

CP9i Electrical safety

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Explain the difference between the functions of the live and the neutral wires.			
4 ^{ch}	Explain how circuit breakers make circuits safer.			
4 th	Explain how the earth wire and the fuse make circuits safer.			
4 th	Explain why switches and fuses are connected in the live wire.			
5 th	Recall the potential differences between the live, neutral and earth wires.			
4 th	Explain the danger of a connection between the live wire and earth.			

CP10 Magnetism and the Motor Effect (Paper 6: Physics 2)

CP10a Magnets and magnetic fields

Step	Learning outcome	Had a look	Nearly there	Nailed it!
3rd	Describe how magnets affect each other.			
4 th	Explain the difference between permanent and induced magnets.			
4 th	Describe the uses of permanent and temporary magnetic materials.			
4 th	Describe the shapes of magnetic fields, including variations in strength.			
4 th	Describe how the shape of magnetic fields can be shown using plotting compasses.			
5 th	Explain how a magnetic compass can be used as evidence for the Earth's magnetic core.			

CP10b Electromagnetism

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall that a current can create a magnetic effect.			
7 th	Relate the shape and direction of the magnetic field around a straight wire to the direction of the current.			
6 th	Recall the factors that affect the strength of the magnetic field around a wire.			
7 th	Describe the magnetic field inside and outside a coil of wire carrying a current.			
8 th	Explain the shape and strength of the magnetic field around a solenoid.			

CP10c Magnetic forces

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall that forces are produced when a current flows in a magnetic field.			
7 th	Explain what causes the forces produced when a current flows in a magnetic field.			
6 th	Recall Fleming's left-hand rule.			
7 th	Use Fleming's left-hand rule.			
8 th	Use the formula relating force, magnetic field strength, current and length.			

CP11 Electromagnetic Induction (Paper 6: Physics 2)

CP11a Transformers

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall the law of conservation of energy.			
	Recall that the power of an electrical current is given by the current multiplied by the voltage.			
	Use the formula relating the input and output current and voltage for a transformer.			

CP11b Transformers and energy

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Recall the factors that affect the size and direction of an induced potential difference.			
7 th	Describe how the magnetic field produced by an induced potential difference opposes the original change.			
8 th	H Explain how a transformer works.			
6 th	Recall that transformers can change the voltage of an alternating current.			
6 th	Describe how the national grid transmits electricity around the country.			
7 th	Explain why step-up and step-down transformers are used in the national grid.			

CP12 Particle Model (Paper 6: Physics 2)

CP12a Particles and density

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Describe the arrangements of particles in solids, liquids and gases.			
7 th	Use the particle model to explain the different properties of solids, liquids and gases.			
5 th	Recall the formula relating density, mass and volume.			
7 th	Use the formula relating density, mass and volume.			
7 th	Use the particle model to explain why solids, liquids and gases have different densities.			
4 th	Describe what happens to the mass of a substance when it changes state.			

CP12b Energy and changes of state

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how heating affects the particles in a substance or object, including changes of state.			
6 th	Describe how the temperature of an object changes with time while being heated or cooled to make it change state.			
6 th	Define the term specific heat capacity.			
6 th	Define the term specific latent heat.			
8 th	Explain the difference between specific heat capacity and specific latent heat.			
6 th	Explain ways of reducing unwanted energy transfer through thermal insulation.			

CP12c Energy calculations

Step	Learning outcome	Had a look	Nearly there	Nailed it!
8 th	Use the formula relating change in thermal energy, mass, temperature change and specific heat capacity.			
8 th	Use the formula relating thermal energy, mass and specific latent heat.			
6 th	Recall that the value of specific latent heat for a substance is different for melting/solidifying and for evaporating/condensing.			

CP12d Gas temperature and pressure

Step	Learning outcome	Had a look	Nearly there	Nailed it!
6 th	Explain how the movement of particles causes gas pressure.			
6 th	Explain how changing the temperature of a gas affects the speed of its particles.			
6 th	Explain how temperature affects the pressure of a fixed mass of gas at constant volume.			
6 th	Explain the significance of absolute zero.			
6 th	Convert temperatures between the Kelvin and Celsius temperature scales.			

CP13 Forces and Matter (Paper 6: Physics 2)

CP13a Bending and stretching

Step	Learning outcome	Had a look	Nearly there	Nailed it!
4 th	Explain that more than one force is needed to distort an object.			
4 th	Describe the difference between elastic and inelastic distortion.			
4 th	Describe the relationship between force and extension for a spring.			
4 th	Describe the relationship between force and extension for a rubber band.			
6 th	Compare the force–extension relationship for different objects.			

CP13b Extension and energy transfers

Step	Learning outcome	Had a look	Nearly there	Nailed it!
5 th	Recall the equation that links force, extension and the spring constant.			
7 th	Use the formula relating force, extension and spring constant.			
5 th	Recall that work has to be done to stretch a spring.			
7 th	Use the formula relating the energy transferred to the extension of a spring.			