

AP Chemistry Summer Assignment

The summer assignment for AP Chemistry is quite simple (but not easy). You need to master the formulas, charges, and names of the common ions. In the first week of the school year, you will be given a quiz on these ions. You will be asked to:

- write the names of these ions when given the formula and charge
- write the formula and charge when given the names

I have included several resources in this packet. First, there is a list of the ions that you must know on the first day. This list also has, on the back, some suggestions for making the process of memorization easier. For instance, many of you will remember that most of the monatomic ions have charges that are directly related to their placement on the periodic table. There are naming patterns that greatly simplify the learning of the polyatomic ions as well.

Also included is a copy of the periodic table used in AP Chemistry. Notice that this is not the table used in first year chemistry. The AP table is the same that the College Board allows you to use on the AP Chemistry test. Notice that it has the symbols of the elements but not the written names. You need to take that fact into consideration when studying for the afore-mentioned quiz!

I have included a sheet of flashcards for the polyatomic ions that you must learn. I strongly suggest that you cut them out and begin memorizing them immediately. Use the hints on the common ions sheet to help you reduce the amount of memorizing that you must do.

Do not let the fact that there are no flashcards for monatomic ions suggest to you that the monatomic ions are not important. They are every bit as important as the polyatomic ions. If you have trouble identifying the charge of monatomic ions (or the naming system) then I suggest that you make yourself some flashcards for those as well.

Doubtless, there will be some students who will procrastinate and try to do all of this studying just before the start of school. Those students may even cram well enough to do well on the initial quiz. However, they will quickly forget the ions, and struggle every time that these formulas are used in lecture, homework, quizzes, tests and labs. All research on human memory shows us that frequent, short periods of study, spread over long periods of time will produce much greater retention than long periods of study of a short period of time.

I could wait and throw these at you on the first day of school, but I don't think that would be fair to you. Use every modality possible as you try to learn these – speak them, write them, visualize them.

I look forward to seeing you all at the beginning of the next school year. If you need to contact me during the summer, you can email me and I will get back to you quickly.

Best of luck to you all,

Ms. Moss

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Common Ions and Their Charges

A mastery of the common ions, their formulas and their charges, is essential to success in AP Chemistry. You are expected to know all of these ions on the first day of class, when I will give you a quiz on them. You will always be allowed a periodic table, which makes identifying the ions on the left "automatic." For tips on learning these ions, see the opposite side of this page.

From the table:	
Cations	Name
H ₁₊	Hydrogen
Li ₁₊	Lithium
Na ₁₊	Sodium
K ₁₊	Potassium
Rb ₁₊	Rubidium
Cs ₁₊	Cesium
Be ₂₊	Beryllium
Mg ₂₊	Magnesium
Ca ₂₊	Calcium
Ba ₂₊	Barium
Sr ₂₊	Strontium
Al ₃₊	Aluminum
Anions	Name
H ₁₋	Hydride
F ₁₋	Fluoride
Cl ₁₋	Chloride
Br ₁₋	Bromide
I ₁₋	Iodide
O ₂₋	Oxide
S ₂₋	Sulfide
Se ₂₋	Selenide
N ₃₋	Nitride
P ₃₋	Phosphide
As ₃₋	Arsenide
Type II Cations	Name
Fe ₃₊	Iron(III) (Ferrous)
Fe ₂₊	Iron(II) (Ferric)
Cu ₂₊	Copper(II) (Cupric)
Cu ₁₊	Copper(I) (Cuprous)
Co ₃₊	Cobalt(III) (Cobaltic)
Co ₂₊	Cobalt(II) (Cobaltous)
Sn ₄₊	Tin(IV) (stannous)
Sn ₂₊	Tin(II) (Stannic)
Pb ₄₊	Lead(IV) (Plumbic)
Pb ₂₊	Lead(II) (Plumbous)
Hg ₂₊	Mercury(II)

Ions to Memorize	
Cations	Name
Ag ₁₊	Silver
Zn ₂₊	Zinc
Hg ₂₊	Mercury(I)
NH ₄₊	Ammonium
Anions	Name
NO ₂₋	Nitrite
NO ₃₋	Nitrate
SO ₃₋	Sulfite
SO ₄₋	Sulfate
HSO ₄₋	Hydrogen sulfate (bisulfate)
OH ₁₋	Hydroxide
CN ₁₋	Cyanide
PO ₄₋	Phosphate
HPO ₄₋	Hydrogen phosphate
H ₂ PO ₄₋	Dihydrogen phosphate
NCS ₁₋	Thiocyanate
CO ₃₋	Carbonate
HCO ₃₋	Hydrogen carbonate (bicarbonate)
ClO ₁₋	Hypochlorite
ClO ₂₋	Chlorite
ClO ₃₋	Chlorate
ClO ₄₋	Perchlorate
BrO ₁₋	Hypobromite
BrO ₂₋	Bromite
BrO ₃₋	Bromate
BrO ₄₋	Perbromate
IO ₁₋	Hypoiodite
IO ₂₋	iodite
IO ₃₋	iodate
IO ₄₋	Periodate
C ₂ H ₃ O ₂₋	Acetate
MnO ₄₋	Permanganate
Cr ₂ O ₇₋	Dichromate
CrO ₄₋	Chromate
O ₂₋	Peroxide

C_2O_{42-}	Oxalate
NH_{21-}	Amide
BO_{33-}	Borate
S_2O_{32-}	Thiosulfate

Tips for Learning the Ions

“From the Table”

These are ions can be organized into two groups.

1. Their place on the table suggests the charge on the ion, since the neutral atom gains or loses a predictable number of electrons in order to obtain a noble gas configuration. This was a focus in first year chemistry, so if you are unsure what this means, get help BEFORE the start of the year.
 - a. All Group 1 Elements (alkali metals) lose one electron to form an ion with a 1+ charge
 - b. All Group 2 Elements (alkaline earth metals) lose two electrons to form an ion with a 2+ charge
 - c. Group 13 metals like aluminum lose three electrons to form an ion with a 3+ charge
 - d. All Group 17 Elements (halogens) gain one electron to form an ion with a 1- charge
 - e. All Group 16 nonmetals gain two electrons to form an ion with a 2- charge
 - f. All Group 15 nonmetals gain three electrons to form an ion with a 3- charge

Notice that cations keep their name (sodium ion, calcium ion) while anions get an “-ide” ending (chloride ion, oxide ion).

2. Metals that can form more than one ion will have their positive charge denoted by a roman numeral in parenthesis immediately next to the name of the

Polyatomic Anions

Most of the work on memorization occurs with these ions, but there are a number of patterns that can greatly reduce the amount of memorizing that one must do.

1. “ate” anions have one more oxygen then the “ite” ion, but the same charge. If you memorize the “ate” ions, then you should be able to derive the formula for the “ite” ion and vice-versa.
 - a. sulfate is SO_{42-} , so sulfite has the same charge but one less oxygen (SO_{32-})
 - b. nitrate is NO_{31-} , so nitrite has the same charge but one less oxygen (NO_{21-})
2. If you know that a sulfate ion is SO_{42-} then to get the formula for hydrogen sulfate ion, you add a hydrogen ion to the front of the formula. Since a hydrogen ion has a 1+ charge, the net charge on the new ion is less negative by one.
 - a. Example:

PO_{43-}	\rightleftharpoons	HPO_{42-}	\rightleftharpoons	H_2PO_{41-}
phosphate		hydrogen phosphate		dihydrogen phosphate
3. Learn the hypochlorite \rightleftharpoons chlorite \rightleftharpoons chlorate \rightleftharpoons perchlorate series, and you also know the series containing iodite/iodate as well as bromite/bromate.
 - a. The relationship between the “ite” and “ate” ion is predictable, as always. Learn one and you know the other.
 - b. The prefix “hypo” means “under” or “too little” (think “hypodermic”, “hypothermic” or “hypoglycemia”)
 - i. Hypochlorite is “under” chlorite, meaning it has one less oxygen
 - c. The prefix “hyper” means “above” or “too much” (think “hyperkinetic”)

- i. the prefix “per” is derived from “hyper” so perchlorate (hyperchlorate) has one more oxygen than chlorate.
- d. Notice how this sequence increases in oxygen while retaining the same charge:



DO NOT DETACH FROM BOOK.

PERIODIC TABLE OF THE ELEMENTS

1 H 1.008																	2 He 4.00
3 Li 6.94	4 Be 9.01											5 B 10.81	6 C 12.01	7 N 14.01	8 O 16.00	9 F 19.00	10 Ne 20.18
11 Na 22.99	12 Mg 24.30											13 Al 26.98	14 Si 28.09	15 P 30.97	16 S 32.06	17 Cl 35.45	18 Ar 39.95
19 K 39.10	20 Ca 40.08	21 Sc 44.96	22 Ti 47.90	23 V 50.94	24 Cr 52.00	25 Mn 54.94	26 Fe 55.85	27 Co 58.93	28 Ni 58.69	29 Cu 63.55	30 Zn 65.39	31 Ga 69.72	32 Ge 72.59	33 As 74.92	34 Se 78.96	35 Br 79.90	36 Kr 83.80
37 Rb 85.47	38 Sr 87.62	39 Y 88.91	40 Zr 91.22	41 Nb 92.91	42 Mo 95.94	43 Tc (98)	44 Ru 101.1	45 Rh 102.91	46 Pd 106.42	47 Ag 107.87	48 Cd 112.41	49 In 114.82	50 Sn 118.71	51 Sb 121.75	52 Te 127.60	53 I 126.91	54 Xe 131.29
55 Cs 132.91	56 Ba 137.33	*57 La 138.91	72 Hf 178.49	73 Ta 180.95	74 W 183.85	75 Re 186.21	76 Os 190.2	77 Ir 192.2	78 Pt 195.08	79 Au 196.97	80 Hg 200.59	81 Tl 204.38	82 Pb 207.2	83 Bi 208.98	84 Po (209)	85 At (210)	86 Rn (222)
87 Fr (223)	88 Ra 226.02	†89 Ac 227.03	104 Rf (261)	105 Db (262)	106 Sg (266)	107 Bh (264)	108 Hs (277)	109 Mt (268)	110 Ds (271)	111 Rg (272)							

*Lanthanide Series

58 Ce 140.12	59 Pr 140.91	60 Nd 144.24	61 Pm (145)	62 Sm 150.4	63 Eu 151.97	64 Gd 157.25	65 Tb 158.93	66 Dy 162.50	67 Ho 164.93	68 Er 167.26	69 Tm 168.93	70 Yb 173.04	71 Lu 174.97
90 Th 232.04	91 Pa 231.04	92 U 238.03	93 Np (237)	94 Pu (244)	95 Am (243)	96 Cm (247)	97 Bk (247)	98 Cf (251)	99 Es (252)	100 Fm (257)	101 Md (258)	102 No (259)	103 Lr (262)

†Actinide Series

INFORMATION IN THE TABLE BELOW AND IN THE TABLES ON PAGES 3-5 MAY BE USEFUL IN ANSWERING THE QUESTIONS IN THIS SECTION OF THE EXAMINATION.

Sulfite	Sulfate	Hydrogen sulfate
Phosphate	Dihydrogen Phosphate	Hydrogen Phosphate
Nitrite	Nitrate	Ammonium
Thiocyanate	Carbonate	Hydrogen carbonate
Borate	Chromate	Dichromate
Permanganate	Oxalate	Amide
Hydroxide	Cyanide	Acetate
Peroxide	Hypochlorite	Chlorite
Chlorate	Perchlorate	Thiosulfate



or



or

